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# JEE

## (Main)

**PAPER-1 (B.E./B. TECH.)**

# 2023

## COMPUTER BASED TEST (CBT) Questions & Solutions

**Date: 08 April, 2023 (SHIFT-1) | TIME : (9.00 a.m. to 12.00 p.m)**

**Duration: 3 Hours | Max. Marks: 300**






**SUBJECT: CHEMISTRY**

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**PART : CHEMISTRY**

61. Given below are two statements:

**Statement – I :** Lithium and Magnesium do not form superoxide

**Statement – II :** The ionic radius of  $\text{Li}^+$  is larger than ionic radius of  $\text{Mg}^{2+}$ .

In the light of the above statements, choose the **most appropriate** answer from the options given below:

- (1) Statement I is incorrect but Statement II is correct
- (2) Statement I is correct but Statement II is incorrect
- (3) Both Statement I and Statement II are correct
- (4) Both Statement I and Statement II are incorrect

**Ans. NTA : (3)**

**Sol.**  $r_{\text{Li}^+} \longrightarrow 0.76 \text{ \AA}$

$r_{\text{Mg}^{+2}} \longrightarrow 0.72 \text{ \AA}$

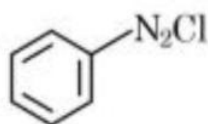
62. What is the purpose of adding gypsum to cement ?

- (1) To facilitate the hydration of cement
- (2) To speed up the process of setting
- (3) To give a hard mass
- (4) To slow down the process of setting

**Ans. NTA : (4)**

**Sol.** The purpose of adding Gypsum is only to slow down the process of setting of the cement.

63. Match List I with List II:



is reacted with reagents in List I to form products in List II.

List I (Reagent)	List II (Product)
A.	I.
B. $\text{HBF}_4, \Delta$	II.
C. $\text{Cu, HCl}$	III.
D. $\text{CuCN / KCN}$	IV.

Choose the correct answer from the options given below:

- (1) A-IV, B-III, C-II, D-I
- (2) A-III, B-I, C-II, D-IV
- (3) A-III, B-I, C-IV, D-II
- (4) A-I, B-III, C-IV, D-II

**Ans. NTA (3)**

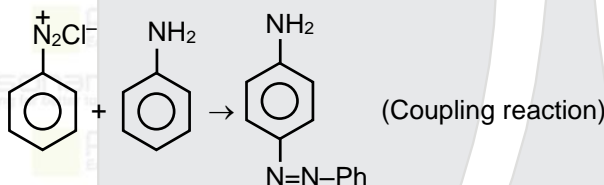
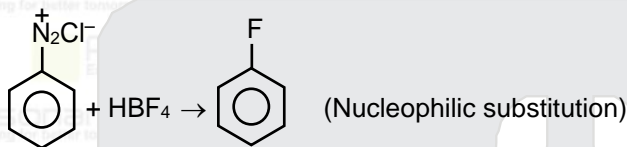
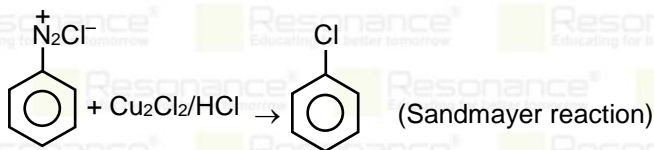
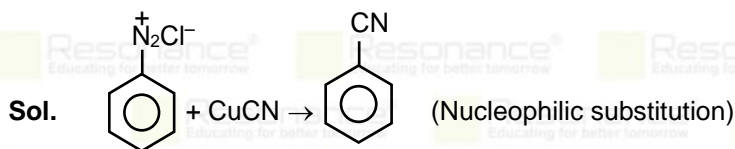
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64. Sulphur (S) containing amino acids from the following are:  
(a) isoleucine (b) cysteine (c) lysine (d) methionine (e) glutamic acid

- (1) b, c, e  
(2) a, b, c  
(3) a, d  
(4) b, d

Ans. NTA (4)

Sol. NCERT based factual question.

65. Which halogen is known to cause the reaction given below :



- (1) Only Chlorine  
(2) Only Bromine  
(3) All halogens  
(4) Only Iodine

NTA (4)

Reso ( )

Sol. CuI<sub>2</sub> is unstable, E<sup>o</sup><sub>cell</sub> for this reaction will be positive.

66. The correct order of spin only magnetic moments for the following complex ions is

- (1) [MnBr<sub>4</sub>]<sup>2-</sup> < [CoF<sub>6</sub>]<sup>3-</sup> < [Fe(CN)<sub>6</sub>]<sup>3-</sup> < [Mn(CN)<sub>6</sub>]<sup>3-</sup>  
(2) [CoF<sub>6</sub>]<sup>3-</sup> < [MnBr<sub>4</sub>]<sup>2-</sup> < [Fe(CN)<sub>6</sub>]<sup>3-</sup> < [Mn(CN)<sub>6</sub>]<sup>3-</sup>  
(3) [Fe(CN)<sub>6</sub>]<sup>3-</sup> < [CoF<sub>6</sub>]<sup>3-</sup> < [MnBr<sub>4</sub>]<sup>2-</sup> < [Mn(CN)<sub>6</sub>]<sup>3-</sup>  
(4) Fe(CN)<sub>6</sub>]<sup>3-</sup> < [Mn(CN)<sub>6</sub>]<sup>3-</sup> < [CoF<sub>6</sub>]<sup>3-</sup> < [MnBr<sub>4</sub>]<sup>2-</sup>

Ans. NTA (4)

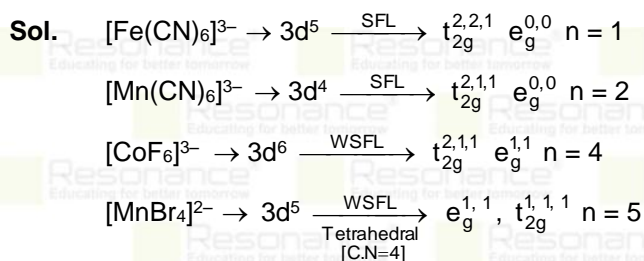
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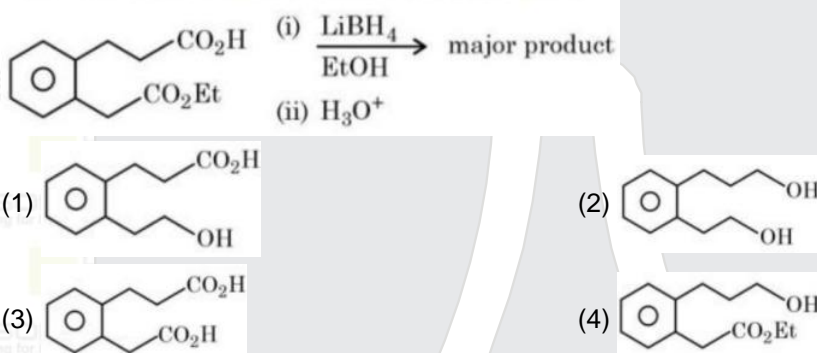
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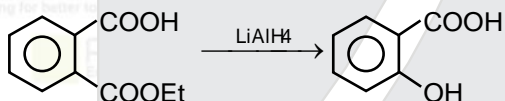


67. The major product formed in the following reaction is:



**Ans.** NTA (1)

**Sol.**  $\text{LiBH}_4 / \text{EtOH}$  is selective reducing agent which reduces carboxylic esters against carboxylic acids reaction is chemo selective.



68. Which of the following metals can be extracted through alkali leaching technique ?

- (1) Au (2) Cu (3) Sn (4) Pb

**Ans.** NTA (3)

**Sol.** Cassiterite ore  $[\text{SnO}_2]$  is Amphoteric in nature. This property is used during hydrometallurgy of cassiterite ore. During alkali leaching  $\text{NaOH}/\text{KOH}$  can be used.

69. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason R. Assertion A: Butan -1 - ol has higher boiling point than ethoxyethane.

Reason R: Extensive hydrogen bonding leads to stronger association of molecules.

In the light of the above statements, choose the correct answer from the options given below:

- (1) A is false but R is true  
 (2) Both A and R are true and R is the correct explanation of A  
 (3) Both A and R are true but R is not the correct explanation of A  
 (4) A is true but R is false

**Ans.** NTA (2)

**Sol.** Due to intermolecular hydrogen bonding in Butane-2-ol, association of molecule increases which increase attraction between molecules and boiling point increase.

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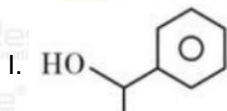
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70. Match List I with List II:

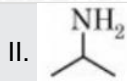
List I (Reagents used)

List II (Compound with Functional group detected)

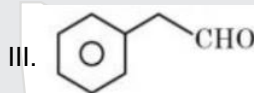
A. Alkaline solution of copper sulphate and sodium citrate



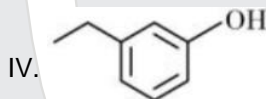
B. Neutral FeCl<sub>3</sub> solution



C. Alkaline chloroform solution



D. Potassium iodide and sodium hypochlorite



Choose the correct answer from the options given below:

- (1) A-III, B-IV, C-I, D-II
- (2) A-IV, B-I, C-II, D-III
- (3) A-II, B-IV, C-III, D-I
- (4) A-III, B-IV, C-II, D-I

Ans. NTA (4)

Sol. CuSO<sub>4</sub> /sodium tartarate → Fehling solution used to identify aldehyde functional group.  
Sodium hypochlorite → Used to find out compound showing haloform test positively.  
Neutral FeCl<sub>3</sub> → Use to identify Phenol & Derivative.  
CHCl<sub>3</sub>/KOH → React with 1° amine and this is known as carbyl amine test.

71. Which of the following complex is octahedral, diamagnetic and the most stable ?

- (1) [Co(H<sub>2</sub>O)<sub>6</sub>]Cl<sub>2</sub>
- (2) K<sub>3</sub>[Co(CN)<sub>6</sub>]
- (3) [Ni(NH<sub>3</sub>)<sub>6</sub>]Cl<sub>2</sub>
- (4) Na<sub>3</sub>[CoCl<sub>6</sub>]

Ans. NTA (2)






Sol. K<sub>3</sub>[Co(CN)<sub>6</sub>] → diamagnetic, d<sup>2</sup>sp<sup>3</sup> & most stable.

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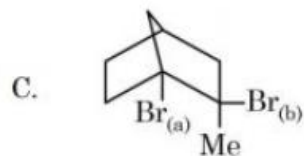
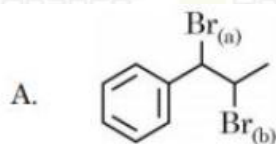
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72. Choose the halogen which is most reactive towards  $S_N1$  reaction in the given compounds (A, B, C & D)



- (1) A – Br(b) ; B – I(b) ; C – Br(b) ; D – Br(b)  
 (2) A – Br(b) ; B – I(a) ; C – Br(a) ; D – Br(a)  
 (3) A – Br(a) ; B – I(a) ; C – Br(a) ; D – Br(a)  
 (4) A – Br(a) ; B – I(a) ; C – Br(b) ; D – Br(a)

Ans. NTA (4)

Sol. The reactivity of halogen for  $S_N1$  reaction depend upon stability of carbocation.

73.  $2IO_3^- + xI^- + 12H^+ \longrightarrow 6I_2 + 6H_2O$

What is the value of x ?

- (1) 6  
 (2) 12  
 (3) 10  
 (4) 2

Ans. NTA (3)

Sol.  $2IO_3^- + 10I^- + 12H^+ \longrightarrow 6I_2 + 6H_2O$

74. The correct order of electronegativity for given elements is :

- (1) P > Br > C > At  
 (2) Br > P > At > C  
 (3) Br > C > At > P  
 (4) C > P > At > Br

Ans. NTA (3)

Sol.






Element	C	P	Br	At
Electronegativity	2.5	2.1	2.8	2.2

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75. Match List I with List II:

List I		List II	
A.	Saccharin	I.	High potency sweetener
B.	Aspartame	II.	First artificial sweetening agent
C.	Alitame	III.	Stable at cooking temperature
D.	Sucralose	IV.	Unstable at cooking temperature

Choose the correct answer from the options given below:

- (1) A-IV, B-III, C-I, D-II
- (2) A-II, B-IV, C-I, D-III
- (3) A-II, B-IV, C-III, D-I
- (4) A-II, B-III, C-IV, D-I

Ans. NTA (2)

Sol. NCERT based factual question.

76. The water gas on reacting with cobalt as a catalyst forms

- (1) Ethanol
- (2) Methanoic acid
- (3) Methanal
- (4) Methanol

Ans. NTA (4)

Sol.  $\text{CO(g)} + 2\text{H}_2\text{(g)} \xrightarrow[\text{Catalyst}]{\text{Cobalt}} \text{CH}_3\text{OH(l)}$

77. Match List I with List II:

List I (Species)		List II (Maximum allowed concentration in ppm in drinking water)	
A.	F <sup>-</sup>	I.	< 50 ppm
B.	SO <sub>4</sub> <sup>2-</sup>	II.	< 5 ppm
C.	NO <sub>3</sub> <sup>-</sup>	III.	< 2 ppm
D.	Zn	IV.	< 500 ppm

Choose the correct answer from the options given below:

- (1) A-I, B-II, C-III, D-IV
- (2) A-IV, B-III, C-II, D-I
- (3) A-II, B-I, C-III, D-IV
- (4) A-III, B-II, C-I, D-IV

Ans. NTA (4)

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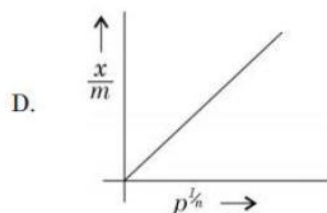
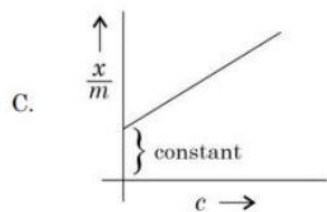
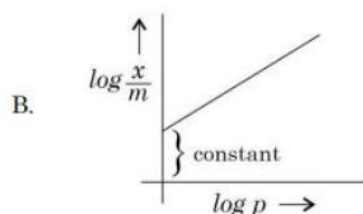
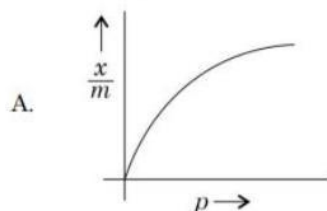
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78. Which of the following represent the Freundlich adsorption isotherms ?



Choose the correct answer from the options given below:

- (1) B, C, D only
- (2) A, B, D only
- (3) A, C, D only
- (4) A, B only

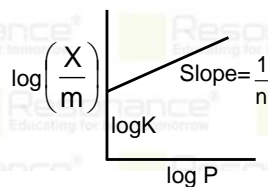
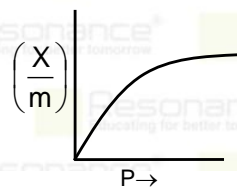
**Ans. NTA (2)**

**Sol.** From Freundlich adsorption isotherm

(a) for adsorption of gas

$$\frac{X}{m} = K(P)^{\frac{1}{n}}$$

$$\log \frac{X}{m} = \log K + \frac{1}{n} \log P.$$



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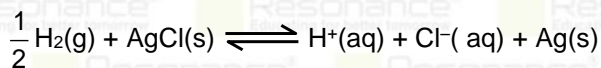
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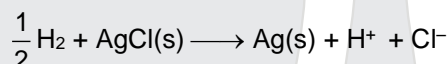
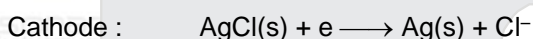
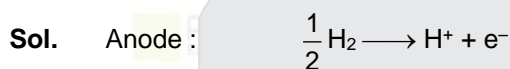
79. The reaction



occurs in which of the given galvanic cell.

- (1) Pt | H<sub>2</sub>(g) | HCl(sol<sup>n</sup>) | AgNO<sub>3</sub>(sol<sup>n</sup>) | Ag
- (2) Pt | H<sub>2</sub>(g) | HCl(sol<sup>n</sup>) | AgCl(s) | Ag
- (3) Ag | AgCl(s) | KCl(sol<sup>n</sup>) | AgNO<sub>3</sub> | Ag
- (4) Pt | H<sub>2</sub>(g) | KCl(sol<sup>n</sup>) | AgCl(s) | Ag

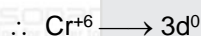
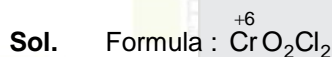
Ans. NTA (2)



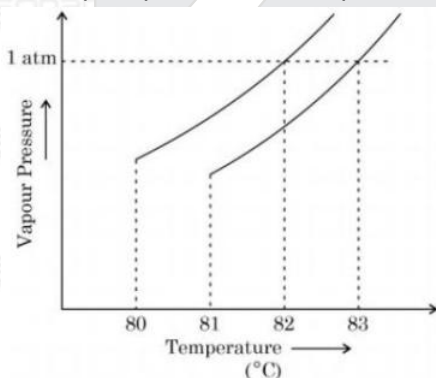
80. In chromyl chloride, the number of d-electrons present on chromium is same as in (Given at no. of Ti : 22, V : 23, Cr : 24, Mn : 25, Fe : 26)

- (1) Mn (VII)
- (2) Fe (III)
- (3) Ti (III)
- (4) V (IV)

Ans. NTA (1)



81. The vapour pressure vs, temperature curve for a solution solvent system is shown below.



The boiling point of the solvent is \_\_\_\_\_ °C.

Given 1.00

Ans. NTA (82)

Sol. the temp. at which V.P. of solvent become 1 atm is called normal B.P. of solvent

Ans. [82°C]

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82. 0.5 g of an organic compound (X) with 60% carbon will produce \_\_\_\_\_ x 10<sup>-2</sup> g of CO<sub>2</sub> on complete combustion.

Ans. NTA (11)

$$\text{Sol. } 60 = \frac{12}{44} \times \frac{\text{wt. of CO}_2}{0.5} \times 100$$

$$\text{wt. of CO}_2 = \frac{60 \times 44 \times 0.5}{12 \times 100} = 1.1$$

83. The number of given statements which is/are correct is \_\_\_\_\_

- (A) The stronger the temperature dependence of the rate constant, the higher is the activation energy.  
 (B) If a reaction has Zero activation energy, its rate is independent of temperature.  
 (C) The stronger the temperature dependence of the rate constant, the smaller is the activation energy.  
 (D) If there is no correlation between the temperature and the rate constant then it means that the reaction has negative activation energy.

Ans. NTA (2)

Reso ( )

$$\text{Sol. } K = A e^{-E_a/RT}$$

$$\frac{1}{K} \cdot \left( \frac{dK}{dT} \right) = \frac{E_a}{RT^2}$$

Ans. (A, B)

84. The number of following factors which affect the percent covalent character of the ionic bond is \_\_\_\_\_

- (A) Polarising power of cation  
 (B) Extent of distortion of anion  
 (C) Polarisability of the anion  
 (D) Polarising power of anion

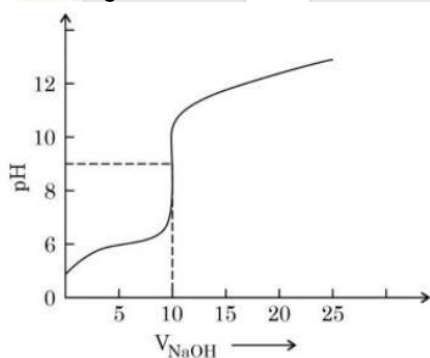
Ans. NTA (3)

Sol. Fajan's rule.

85. The titration curve of weak acid vs. strong base with phenolphthalein as indicator is shown below. The

$$K_{\text{phenolphthalein}} = 4 \times 10^{-10}$$

Given : log 2 = 0.3



The number of following statements which is/are correct about phenolphthalein is \_\_\_\_\_

- A. It can be used as an indicator for the titration of weak acid with weak base.  
 B. It begins to change colour at pH = 8.4  
 C. It is a weak organic base  
 D. It is colourless in acidic medium

Ans. NTA (2)

Sol. 1. Working pH range is 8 to 10

2. It is colourless in acidic medium and pink in basic medium.

3. It is an acidic indicator.

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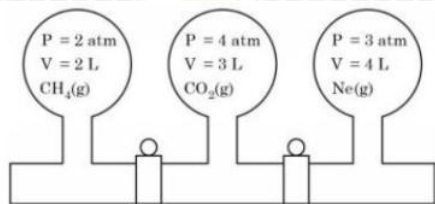
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86.



Three bulbs are filled with  $\text{CH}_4$ ,  $\text{CO}_2$  and  $\text{Ne}$  as shown in the picture. The bulbs are connected through pipes of zero volume. When the stopcocks are opened and the temperature is kept constant throughout, the pressure of the system is found to be \_\_\_\_\_ atm. (Nearest integer)

Ans. NTA (3)

Sol.  $n_1 + n_2 + n_3 = n_{\text{Total}}$


$$\frac{P_1 V_1}{RT} + \frac{P_2 V_2}{RT} + \frac{P_3 V_3}{RT} = \frac{P_T V_T}{RT}$$

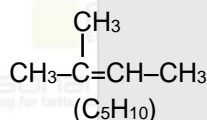
$$2 \times 2 + 4 \times 3 + 3 \times 4 = P_T \times 9$$

$$P_T = 3 \text{ atm}$$

87. Molar mass of the hydrocarbon (X) which on ozonolysis consumes one mole of  $\text{O}_3$  per mole of (X) and gives one mole each of ethanol and propanone is \_\_\_\_\_  $\text{g mol}^{-1}$  (Molar mass of C :  $12 \text{ g mol}^{-1}$ , H :  $1 \text{ g mol}^{-1}$ )

Ans. NTA (70)

Sol. Reactant  $\xrightarrow[\text{Zn-H}_2\text{O}]{\text{O}_3}$   +  $\text{CH}_3\text{CHO}$



88. The number of following statements which is/are incorrect is \_\_\_\_\_

- (A) Line emission spectra are used to study the electronic structure
- (B) The emission spectra of atoms in the gas phase show a continuous spread of wavelength from red to violet
- (C) An absorption spectrum is like the photographic negative of an emission spectrum
- (D) The element helium was discovered in the sun by spectroscopic method

Ans. NTA (1)

Sol. Emission spectrum of an atom is discrete spectrum.

89. When a 60 W electric heater is immersed in a gas for 100s in a constant volume container with adiabatic walls, the temperature of the gas rises by  $5^\circ\text{C}$ . The heat capacity of the given gas is \_\_\_\_\_  $\text{JK}^{-1}$  (Nearest integer)

Ans. NTA (1200)

Sol.  $q = C\Delta T$

$$60 \times 100 = C \times 5$$





$$C = 1200 \text{ JK}^{-1}$$

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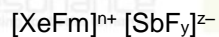
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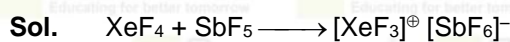
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90. XeF<sub>4</sub> reacts with SbF<sub>5</sub> to form



$$m + n + y + z = \underline{\hspace{2cm}}$$

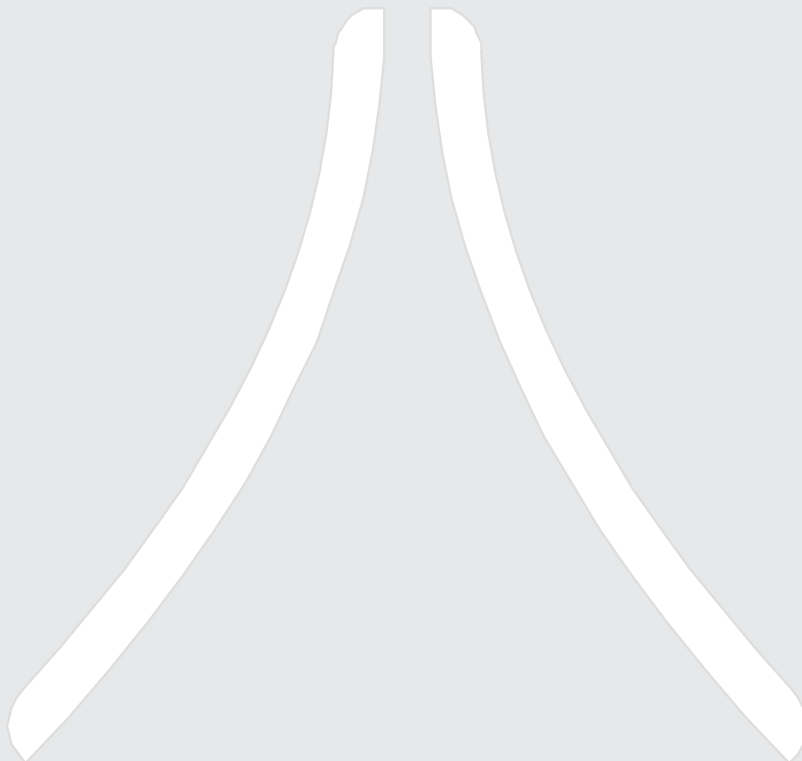
Ans. NTA (11)



$$m = 3 \quad p = 6$$

$$n = 1 \quad q = +1$$

$$m + n + p + q = 3 + 1 + 6 + 1 = 11$$








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