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PAPER-1 (B.E./B. TECH.)

2023

COMPUTER BASED TEST (CBT) Questions & Solutions

Date: 11 April, 2023 (SHIFT-2) | TIME : (3.00 p.m. to 6.00 p.m)

Duration: 3 Hours | Max. Marks: 300






SUBJECT: CHEMISTRY

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PART : CHEMISTRY

61. Alkali metal from the following with least melting point is:

- (1) Rb (2) Na (3) K (4) Cs

Ans. NTA (4)

Sol. On moving down the group size of the kernel increases so metallic bond strength decreases and hence melting point decreases.

62. Which hydride among the following is less stable?

- (1) HF
(2) BeH₂
(3) NH₃
(4) LiH

Ans. NTA (2)

Sol. BeH₂ exist in Polymeric form (BeH₂)_n and Polymerisation occurs through Banana bonding

	3c – 2e ⁻	sp ³ -s-sp ³	Non planar	
(BeH ₂) _n	All bonds except the extreme ends are 3C – 2e ⁻ bonds.			

63. The magnetic moment is measured in Bohr Magnetron (BM).

Spin only magnetic moment of Fe in [Fe(H₂O)₆]³⁺ and [Fe(CN)₆]³⁻ complexes respectively is:

- (1) 5.92 B.M. and 1.732 B.M.
(2) 6.92 B.M. in both
(3) 3.87 B.M. and 1.732 B.M.
(4) 4.89 B.M. and 6.92 B.M.

Ans. NTA (1)

Sol. [Fe(H₂O)₆]³⁺

Oxidation number of Fe = +3

Fe³⁺ = 3d⁵ ; H₂O is W.F.L so configuration is t_{2g}^{1,1,1} e_g^{1,1}

Number of unpaired electron n = 5

magnetic moment (spin only) $\mu = \sqrt{n(n+2)}$ B.M.

so $\mu = \sqrt{35}$ B.M. = 5.85 B.M.

[Fe(CN)₆]³⁻

Oxidation number of Fe = +3

Fe³⁺ = 3d⁵ ; CN⁻ is S.F.L so configuration is t_{2g}^{2,2,1} e_g^{0,0}

Number of unpaired electron n = 1

magnetic moment (spin only) $\mu = \sqrt{n(n+2)}$ B.M.

so $\mu = \sqrt{3}$ B.M. = 1.73 B.M.

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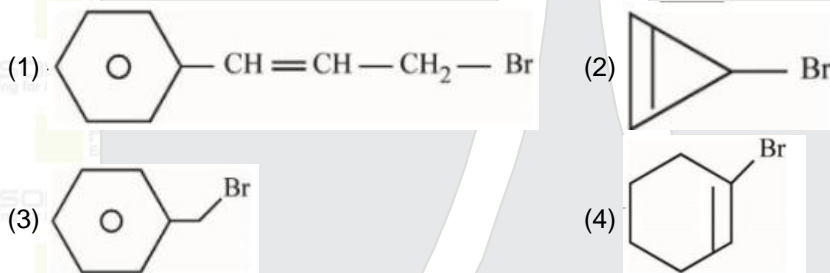
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64. A solution is prepared by adding 2 g of "X" to 1 mole of water. Mass percent of "X" in the solution is
 (1) 20%
 (2) 5%
 (3) 10%
 (4) 2%

Ans. NTA (3)

Sol. Mass % of A in solution = $\frac{W_A}{S_{\text{Solution}}} \times 100$
 $= \frac{2}{20} \times 100 = 10\%$

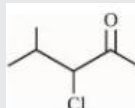
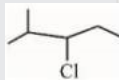
65. Compound from the following that will not produce precipitate on reaction with Compound from the following that will not produce precipitate on reaction with AgNO_3 is: AgNO_3 is:





Ans. NTA (4)

Sol. Vinylic halides are inert for AgNO_3 due to strong C-X bond & less stable carbocation.

66. Given below are two statements. one is labelled as **Assertion A** and the other is labelled as **Reason R**.

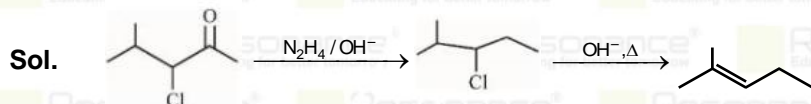
Assertion A:  can be subjected to Wolff-Kishner reduction to give 

Reason R : Wolff-Kishner reduction is used to convert  into 

In the light of the above statements, choose the correct answer from the options given below:

- (1) A is true but R is false
 (2) Both A and R are true and R is the correct explanation of A
 (3) A is false but R is true
 (4) Both A and R are true but R is NOT the correct explanation of A

Ans. NTA (3)








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67. For a chemical reaction $A + B \rightarrow$ Product. the order is 1 with respect to A and B.

Rate $\text{mol L}^{-1} \text{S}^{-1}$	[A] mol L^{-1}	[B] mol L^{-1}
0.10	20	0.5
0.40	x	0.5
0.80	40	y

What is the value of x and y?

- (1) 80 and 4
- (2) 160 and 4
- (3) 40 and 4
- (4) 80 and 2

Ans. NTA (4)

Sol. For reaction $A + B \rightarrow$ product ; rate = $k[A]^1[B]^1$

From Exp. No. 2/1

$$\frac{0.4}{0.1} = \frac{k[x]^1[0.5]^1}{k[20]^1[0.5]^1} ; \frac{4}{1} = \frac{x}{20} ; x = 80$$

From Exp. No. 3/1

$$\frac{0.8}{0.1} = \frac{k[40]^1[y]^1}{k[20]^1[0.5]^1} ; \frac{8}{1} = \frac{40 \times y}{20 \times 0.5} ; y = 2$$

68. One mole of P_4 reacts with 8 moles of SOCl_2 to give 4 moles of A, x mole of SO_2 and 2 moles of B. A, B and x respectively are

- (1) POCl_3 , S_2Cl_2 and 4
- (2) PCl_3 , S_2Cl_2 and 2
- (3) PCl_3 , S_2Cl_2 and 4
- (4) POCl_3 , S_2Cl_2 and 2

Ans. NTA (3)

Sol. $P_4 + 8\text{SOCl}_2 \rightarrow 2\text{S}_2\text{Cl}_2 + 4\text{SO}_2 + 4\text{PCl}_3$

69. Given below are two statements. one is labelled as **Assertion A** and the other is labelled as **Reason R**.

Assertion A : A solution of the product obtained by heating a mole of glycine with a mole of chlorine in presence of red phosphorous generates chiral carbon atom.

Reason R : A molecule with 2 chiral carbons is always optically active.

In the light of the above statements, choose the correct answer from the options given below:

- (1) A is true but R is false
- (2) Both A and R are true but R is NOT the correct explanation of A
- (3) Both A and R are true and R is the correct explanation of A
- (4) A is false but R is true

Ans. NTA (1)

Sol. $\text{NH}_2\text{-CH}_2\text{-COOH} \xrightarrow{\text{Cl}_2/\text{redP}} \text{NH}_2\text{-}\overset{*}{\underset{\text{Cl}}{\text{C}}}\text{-COOH}$

Optical activity does not depend on chiral carbon.

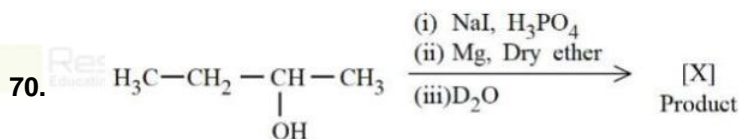
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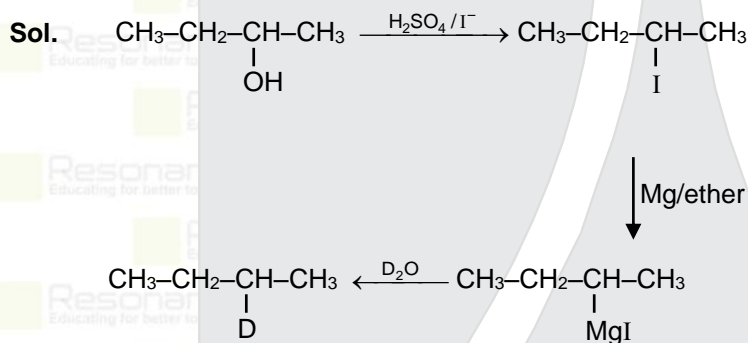
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Product [X] formed in the above reaction is:

- (1)
$$\text{H}_3\text{C}-\text{CH}_2-\underset{\text{OH}}{\overset{\text{H}}{\text{C}}}-\text{CH}_3$$
- (2) $\text{H}_3\text{C}-\text{CH}_2-\text{CH}=\text{CH}_2$
- (3) $\text{H}_3\text{C}-\text{CH}=\text{CH}-\text{CH}_3$
- (4)
$$\text{H}_3\text{C}-\text{CH}_2-\underset{\text{D}}{\text{CH}}-\text{CH}_3$$

Ans. NTA (4)



71. Which of the following compounds is an example of Freon?

- (1) $\text{C}_2\text{Cl}_2\text{F}_2$
- (2) C_2HF_3
- (3) C_2F_4
- (4) $\text{C}_2\text{H}_2\text{F}_2$

Ans. NTA (1)

Sol. $\text{C}_2\text{Cl}_2\text{F}_2$ is 012 Freon

72. Given below are two statements:

Statement I : Ethene at 333 to 343 K and 6-7 atm pressure in the presence of AlEt_3 and TiCl_4 undergoes addition polymerization to give LDP.

Statement II: Caprolactam at 533-543 K in H_2O through step growth polymerizes to give Nylon 6.

In the light of the above statements, choose the correct answer from the options given below:

- (1) Both Statement I and Statement II are false
- (2) Statement I is true but Statement II is false
- (3) Both Statement I and Statement II are true
- (4) Statement I is false but Statement II is true

Ans. NTA (4)

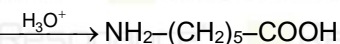
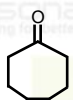
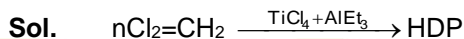
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polymerisation nylon-6

73. If Ni^{2-} is replaced by Pt^{2-} in the complex $[\text{NiCl}_2\text{Br}_2]^{2-}$, which of the following properties are expected to get changed?

- A. Geometry
- B. Geometrical isomerism
- C. Optical isomerism
- D. Magnetic properties

- (1) A, B and C
- (2) B and D
- (3) B and C
- (4) A, B and D

Ans. NTA (4)

Sol. $[\text{NiCl}_2\text{Br}_2]^{2-} \longrightarrow sp^3$ unpaired electron = 2 (No geometrical isomerism)

$[\text{PtCl}_2\text{Br}_2]^{2-} \longrightarrow dsp^2 \longrightarrow$ shows (geometrical isomerism) unpaired electron = 0

No optical isomerism by both.

74. Match List I with List II

LIST I Complex		LIST II Colour	
A.	$\text{Mg}(\text{NH}_4)\text{PO}_4$	I.	brown
B.	$\text{K}_3[\text{Co}(\text{NO}_2)_6]$	II.	white
C.	$\text{MnO}(\text{OH})_2$	III.	yellow
D.	$\text{Fe}_4[\text{Fe}(\text{CN})_6]_3$	IV.	blue

Choose the correct answer from the options given below:

- (1) A-II, B-IV, C-I, D-III
- (2) A-II, B-III, C-I, D-IV
- (3) A-III, B-IV, C-II, D-I
- (4) A-II, B-III, C-IV, D-I

Ans. NTA (2)

Sol. Factual

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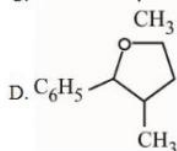
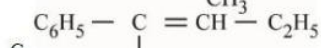
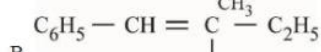
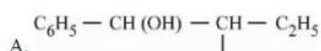
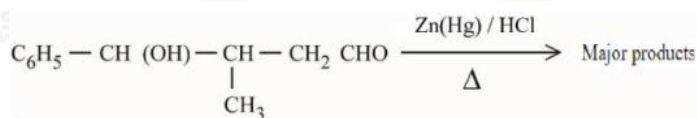
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75. The major product formed in the following reaction is

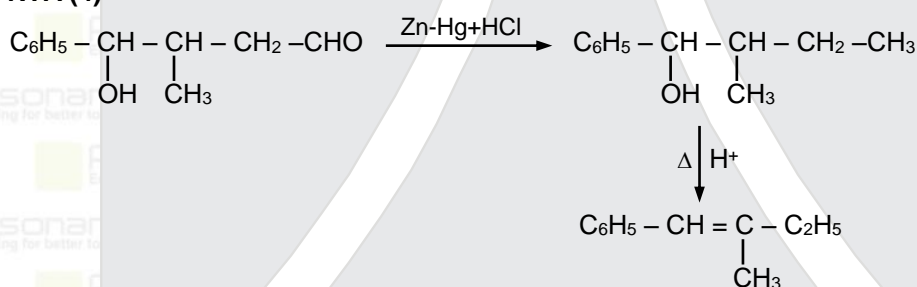


Choose the correct answer from the options given below:

- (1) C only
- (2) A only
- (3) D only
- (4) B only

Ans. NTA (4)

Sol.



76. Which one of the following pairs is an example of polar molecular solids ?

- (1) HCl (s), AlN (s)
- (2) MgO (s), SO₂(s)
- (3) SO₂(s), NH₃(s)
- (4) SO₂(s), CO₂(s)

Ans. NTA (3)

Sol. Polar molecules are SO₂, NH₃, AlN

Non-polar molecules is CO₂

77. Given below are two statements:

Statement I : In the metallurgy process, sulphide ore is converted to oxide before reduction.

Statement II : Oxide ores in general are easier to reduce.

In the light of the above statements, choose the most appropriate answer from the options given below:

- (1) Both Statement I and Statement II are correct
- (2) Statement I is correct but Statement II is incorrect
- (3) Statement I is incorrect but Statement II is correct
- (4) Both Statement I and Statement II are incorrect






Ans. NTA (1)

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Sol. $MS \xrightarrow{O_2} MO + SO_2 \uparrow$
oxide can be easily reduced.

78. What weight of glucose must be dissolved in 100 g of water to lower the vapour pressure by 0.20 mm Hg?

(Assume dilute solution is being formed)

Given : Vapour pressure of pure water is 54.2 mm Hg at room temperature. Molar mass of glucose is 180 g mol^{-1} .

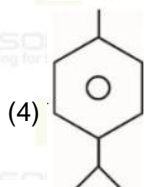
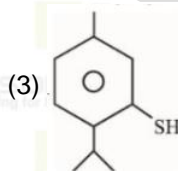
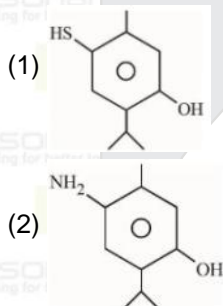
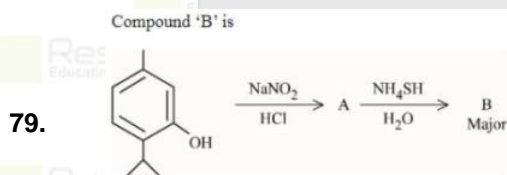
- (1) 4.69 g
(2) 2.59 g
(3) 3.59 g
(4) 3.69 g

Ans. NTA (4)

Sol. For a solution $\frac{P_A^0 - P_s}{P_s} = \frac{n}{N}$

$$\frac{54.2 - 54}{54} = \frac{w/180}{100/18}$$

$$\text{Amount of glucose } w = \frac{200}{54} = 3.69 \text{ g}$$



Ans. NTA (2)

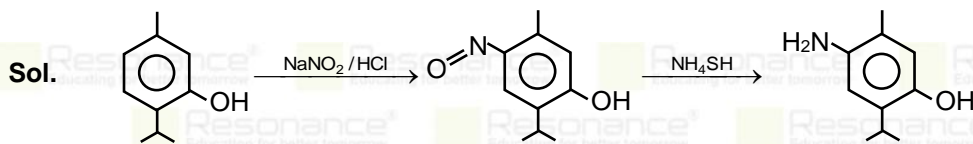
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80. Given below are two statements, one is labelled as **Assertion A** and the other is labelled as Reason R.
Assertion A: $[\text{CoCl}(\text{NH}_3)]^{2-}$ absorbs at lower wavelength of light with respect to $[\text{Co}(\text{NH}_3)_5(\text{H}_2\text{O})]^{3-}$
Reason R: It is because the wavelength of the light absorbed depends on the oxidation state of the metal ion.

in the light of the above statements, choose the correct answer from the options given below:

- (1) Both A and R are true and R is the correct explanation of A
 (2) A is true but R is false
 (3) Both A and R are true but R is NOT the correct explanation of A
 (4) A is false but R is true

Ans. **NTA (4)**

Sol. Cl^- is weaker than water

$$\Delta \quad \text{Cl}^- < \text{H}_2\text{O}$$

$$\lambda \quad \text{Cl}^- > \text{H}_2\text{O}$$

oxidation state affects value of

Δ as well as λ

81. $\text{Mg}(\text{NO}_3)_2 \cdot X\text{H}_2\text{O}$ and $\text{Ba}(\text{NO}_3)_2 \cdot Y\text{H}_2\text{O}$, represent formula of the crystalline forms of nitrate salts. Sum of X and Y is _____.

Ans. **NTA (6)**

Sol. **Factual**

82. The total number of intensive properties from the following is _____

Volume, Molar heat capacity, Molarity, E^\ominus cell, Gibbs free energy change, Molar mass, Mole

Ans. **NTA (4)**

Sol. Intensive properties does not depends of mass or size of the system or substance

Molar Gibb's Free energy, Molar mass, Molarity & specific heat capacity are intensive properties

83. The maximum number of lone pairs of electron on the central atom from the following species is _____

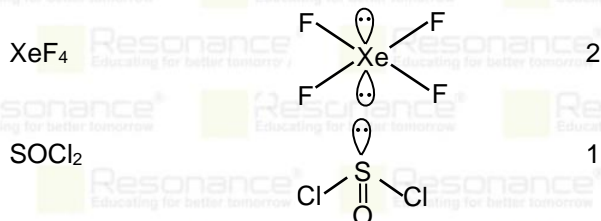
ClO_3^- , XeF_4 , SF_4 and I_3^-

Ans. **NTA (3)**

Sol. **Molecule/Species**

Structure

No. of lone pairs on central atom








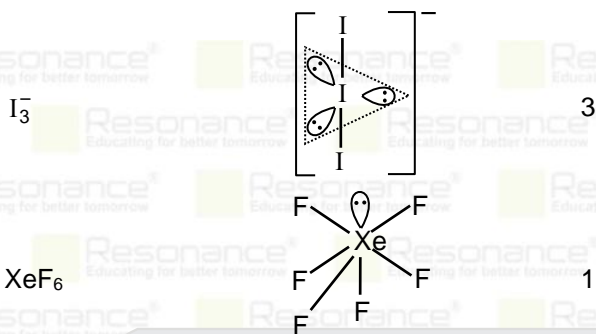
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84. The number of correct statements from the following is _____
- E_{cell} is an intensive parameter
 - A negative E^\ominus means that the redox couple is a stronger reducing agent than the H^+ / H_2 couple.
 - The amount of electricity required for oxidation or reduction depends on the stoichiometry of the electrode reaction.
 - The amount of chemical reaction which occurs at any electrode during electrolysis by a current is proportional to the quantity of electricity passed through the electrolyte.

Ans. NTA (4)

Sol. E_{cell} is intensive property
amount \propto charge (Faraday's law)

85. The number of correct statements about modern adsorption theory of heterogeneous catalysis from the following is _____
- The catalyst is diffused over the surface of reactants.
 - Reactants are adsorbed on the surface of the catalyst.
 - Occurrence of chemical reaction on the catalysts surface through formation of an intermediate.
 - It is a combination of intermediate compound formation theory and the old adsorption theory
 - It explains the action of the catalyst as well as those of catalytic promoters and poisons.

Ans. NTA (3)

Sol. Reactants are adsorbed on surface of catalyst intermediate are formed.

86. Number of compounds from the following which will not produce orange red precipitate with Benedict solution is _____
- Glucose, maltose, sucrose, ribose, 2-deoxyribose, amylose, lactose

Ans. NTA ans (3), Reso ans. (2).

Sol. Sucrose & amylose are non-reducing carbohydrates

87. The number of correct statement from the following is _____ .
- For 1s orbital, the probability density is maximum at the nucleus
 - For 2s orbital, the probability density first increases to maximum and then decreases sharply to zero.
 - Boundary surface diagram of the orbitals encloses a region of 100% probability of finding the electron.
 - p and d-orbitals have 1 and 2 angular nodes respectively
 - probability density of p-orbital is zero at the nucleus

Ans. NTA (3)

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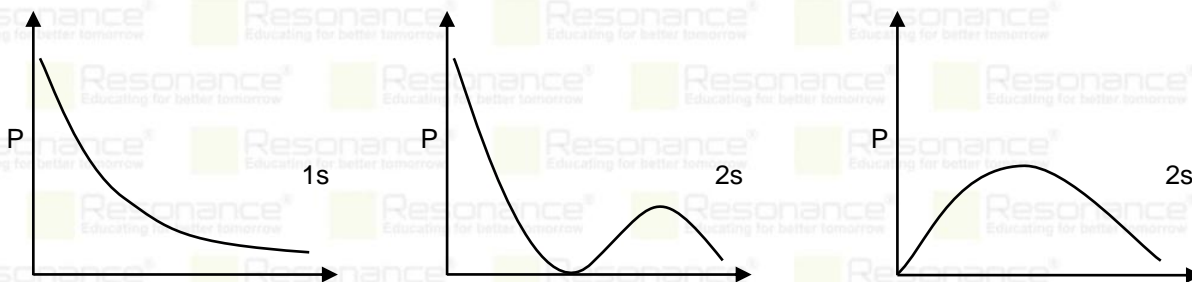
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Sol. P has 1 angular node & d has 2 angular node



88. The volume of hydrogen liberated at STP by treating 2.4 g of magnesium with excess of hydrochloric acid is _____ $\times 10^{-2}$ L

Given : Molar volume of gas is 22.4 L at STP.

Molar mass of magnesium is 24 g mol⁻¹

Ans. NTA (224)

Sol. $\text{Mg} + 2\text{HCl} \longrightarrow \text{MgCl}_2 + \text{H}_2$

1 mole 1 mole

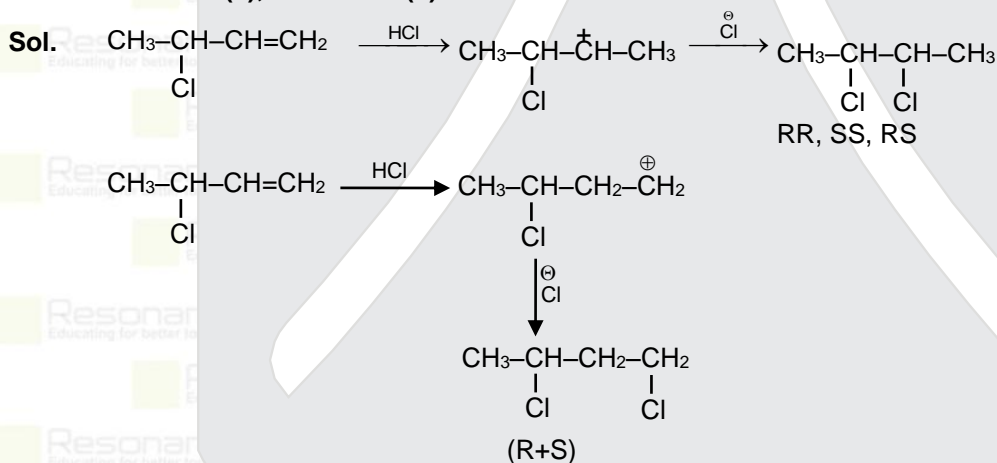
2.4 g of Mg

= 0.1 moles 0.1 moles

So volume of H₂ gas evolved at STP = 0.1 \times 22.4 = 2.24 litre

89. The number of possible isomeric products formed when 3-chloro-1-butene reacts with HCl through carbocation formation is _____

Ans. NTA Ans (4), Reso ans. (5)



90. 4.5 moles each of hydrogen and iodine is heated in a sealed ten litre vessel. At equilibrium, 3 moles of HI were found. The equilibrium constant for $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$ is _____.

Ans. NTA (1)

Sol. $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$

Initial moles 4.5 4.5 0

Moles at eq. 4.5 - 1.5 4.5 - 1.5 3

=3 = 3

$$K_{\text{equilibrium}} = \frac{[\text{HI}]^2}{[\text{H}_2]^1[\text{I}_2]^1} = \frac{[3]^2}{3 \times 3} = 1$$

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