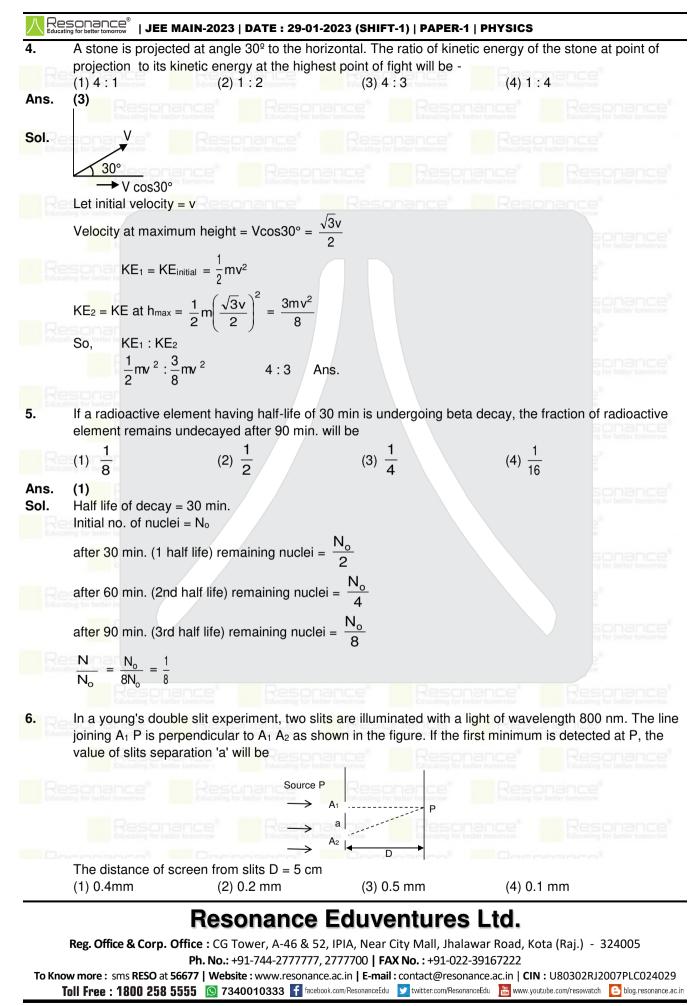
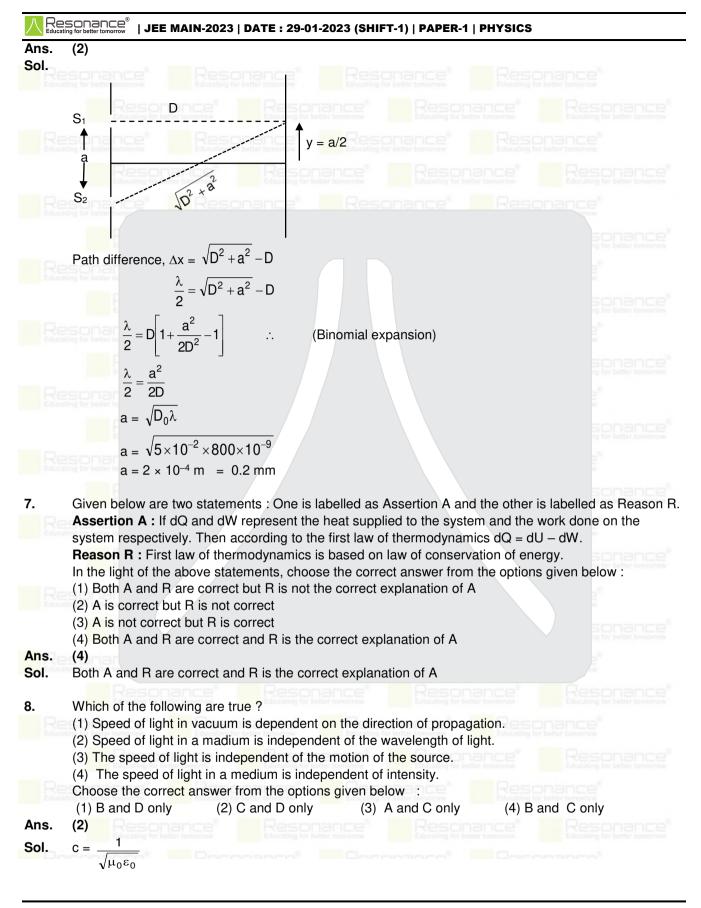


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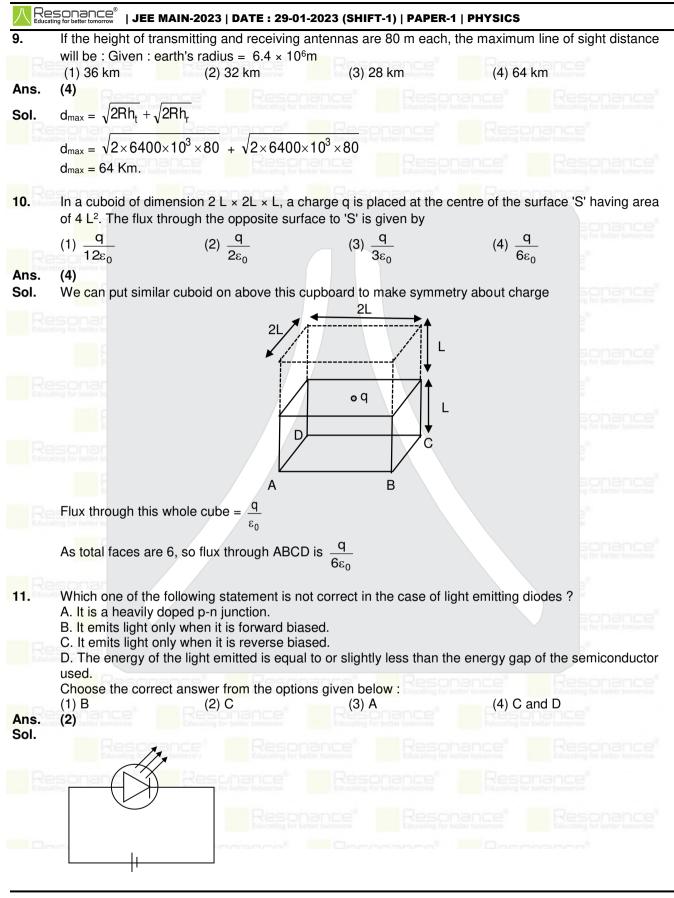
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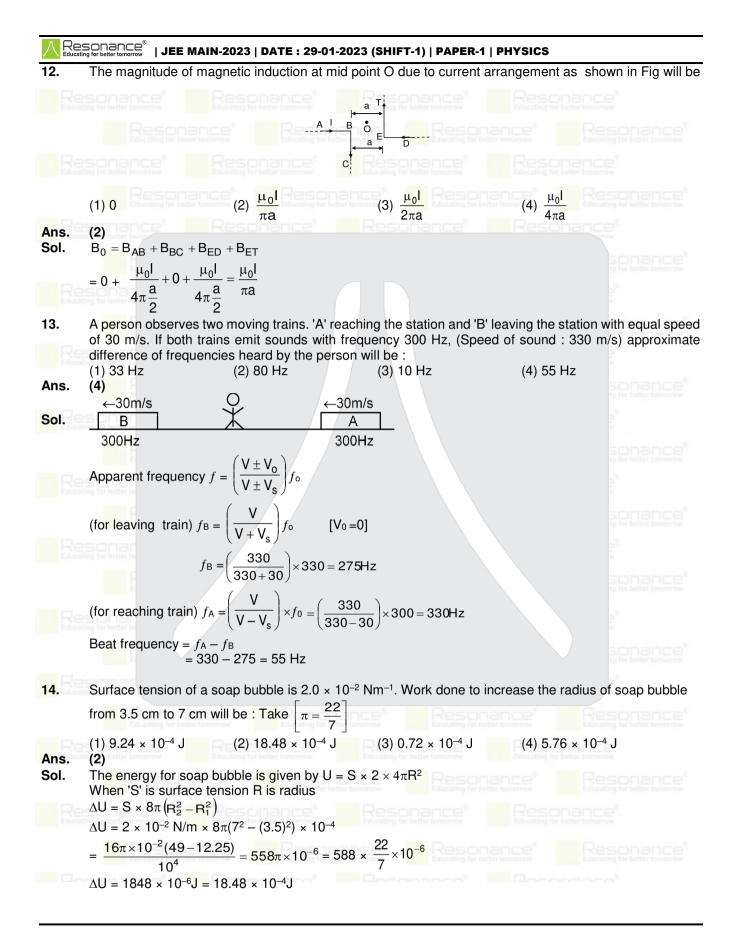
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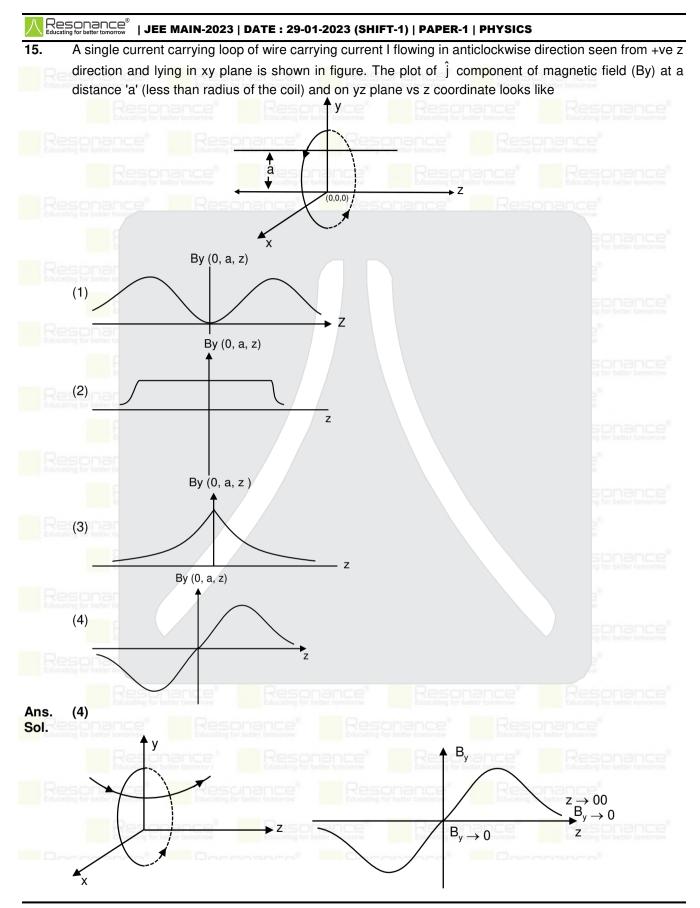
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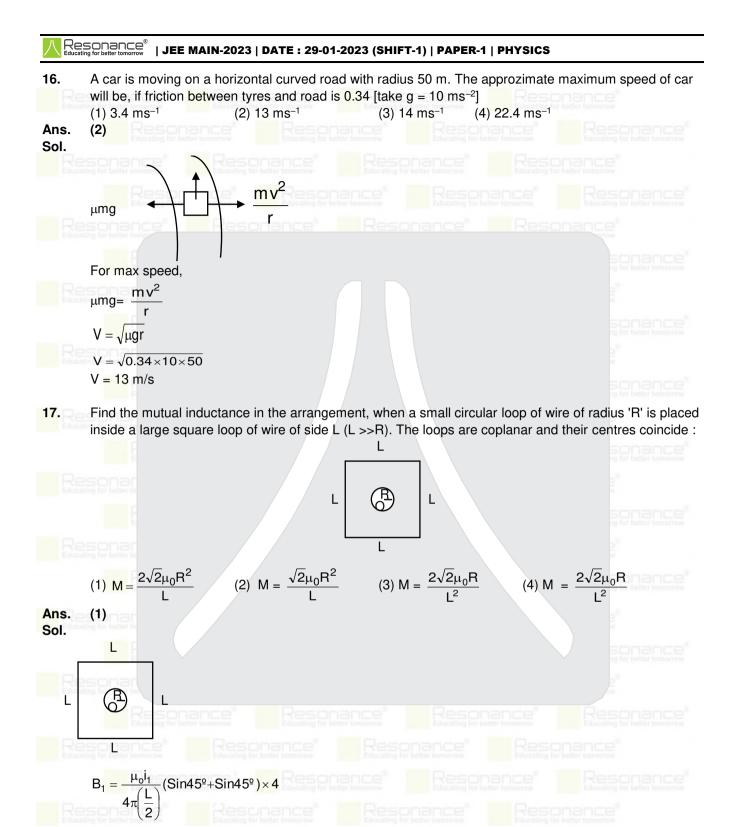
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 $B_{1} = \frac{2\sqrt{2}\mu_{0}i_{1}}{\pi L} ; \qquad \phi_{21} = \left(\frac{2\sqrt{2}\mu_{0}i_{1}}{\pi L}\right)(\pi R^{2}) ; \qquad \phi_{21} = \left(\frac{2\sqrt{2}\mu_{0}\pi R^{2}}{\pi L}\right)i_{1}$ $M_{21} = \frac{2\sqrt{2}\mu_{0}R^{2}}{L}$

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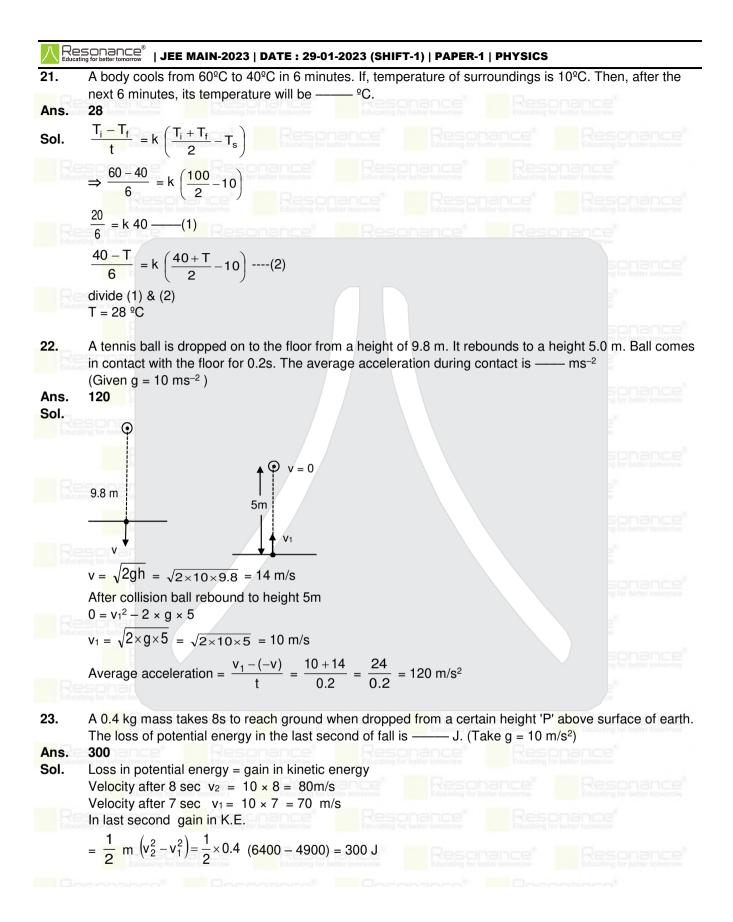
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8.	The threshold wavelength for photoelectric e	emission from a materia	l is 5500 Å. Photoe	electrons will be
	emitted, when this material is illuminated with monochromatic radiation from a			
		(B) 10 W infra-red I		
	(C) 75 W ultra-violet lamp (1 <mark>) C o</mark> nly (2) B and C only	(D) 10W ultra-violet	lamp	
	(1) C only (2) B and C only	(3) C and D only	(4) A and D o	nly
ns.				
iol.	λ_{th} = 550nm, So to eject electrons, the wavelength of light should be $\lambda < \lambda_{th}$			
-	$\lambda < 550$ nm. So ultra-violet light will be able to eject the electrons.			
9.	Match List I with List -II	Peconance'		
Educ	List-I (Physical Quantity)	List-II (Dimensi	onal Formula)	TOW .
	A. Pressure gradient	I. [M ⁰ L ² T ⁻²]		sonance
	B. Energy density	II. [M ¹ L ⁻¹ T ⁻²]		ng for hotter tenterion
	C. Electric Field	III. [M ¹ L ⁻² T ⁻²]		8
	D. Latent heat	I. [M ¹ L ¹ T ⁻² A ⁻¹]		W
	Choose the correct answer from the options			sonance
	(1) A-II, B-III, C-I, D-IV (2) A-III, B-II, C-IV,	0	D-I (4) A-III B-II	C-L D-IV
ns.	(1) / (11, 2) (11, 0) (1, 2) / (11, 2) (1, 0) (1, 1, 0) (1, 1, 0) (1, 1, 1, 0) (1, 1, 1, 0) (1, 1, 1, 0) (1, 1, 1, 1, 0) (1, 1, 1, 1, 1, 1, 1, 1, 1, 1, 1, 1, 1, 1		- · (.), / · · · , D II	
ol.	A-III $\frac{dP}{dx} = \frac{f}{A.L}$			
	$=\frac{MLT^{-2}}{1^{3}}=M^{1}L^{-2}T^{-2}$			
	= Energy MI^2T^{-2}			
	$B-II U_{d} = \frac{Energy}{volume} = \frac{ML^{2}T^{-2}}{L^{3}} = ML^{-1}T^{-2}$			
	aling for better to			
	$C-IV = \frac{F}{q} = \frac{MLT^{-2}}{AT} = M^{1}L^{1}T^{-3}A^{-1}$			
	esonar ^q Al			
	$D-LL = \frac{Heat}{mass} = \frac{energy}{mass} = \frac{ML^2T^{-2}}{M} = L^2T^{-2}$			
	$D = 1 L = \frac{1}{\text{mass}} = \frac{1}{\text{mass}} = \frac{1}{\text{M}} = L^{-1}$			
20.	A bicycle tyre is filled with air having pressure	e of 270 kPa at 27ºC. Tl	ne approximate pre	essure of the a
2 0 . R	A bicycle tyre is filled with air having pressure in the tyre when the temperature increases	e of 270 kPa at 27ºC. TI	ne approximate pre	essure of the a
0. Re		e of 270 kPa at 27ºC. TI (3) 360 kPa	ne approximate pre (4) 278 kPa	essure of the a
20. Ans.	in the tyre when the temperature increases			essure of the a
	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4)			essure of the a
ans.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa			essure of the a
ans.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation)	(3) 360 kPa	(4) 278 kPa	essure of the a
ns.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation) For constant volume	(3) 360 kPa	(4) 278 kPa	essure of the a
ns.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation) For constant volume	(3) 360 kPa	(4) 278 kPa	essure of the a
ns.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation) For constant volume	(3) 360 kPa	(4) 278 kPa	essure of the a
ns.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation) For constant volume $P \propto T$ $\frac{P_1}{P_2} = \frac{T_1}{T_2}$	(3) 360 kPa	(4) 278 kPa	essure of the a
ns.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation) For constant volume $P \propto T$ $\frac{P_1}{P_2} = \frac{T_1}{T_2}$ $\frac{270 \text{KPa}}{T_2} = \frac{(273+27)\text{K}}{T_2}$	(3) 360 kPa	(4) 278 kPa	
ns.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation) For constant volume $P \propto T$ $\frac{P_1}{P_2} = \frac{T_1}{T_2}$ $\frac{270 \text{KPa}}{P_2} = \frac{(273 + 27)\text{K}}{(273 + 36)\text{K}}$	(3) 360 kPa	(4) 278 kPa	essure of the a
ns.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation) For constant volume $P \propto T$ $\frac{P_1}{P_2} = \frac{T_1}{T_2}$ $\frac{270 \text{KPa}}{P_2} = \frac{(273 + 27)\text{K}}{(273 + 36)\text{K}}$	(3) 360 kPa	(4) 278 kPa	
ns.	in the tyre when the temperature increases (1) 262 kPa (2) 270 kPa (4) PV = nRT (ideal gas equation) For constant volume $P \propto T$ $\frac{P_1}{P_2} = \frac{T_1}{T_2}$ $\frac{270 \text{KPa}}{T_2} = \frac{(273+27)\text{K}}{T_2}$	(3) 360 kPa	(4) 278 kPa	

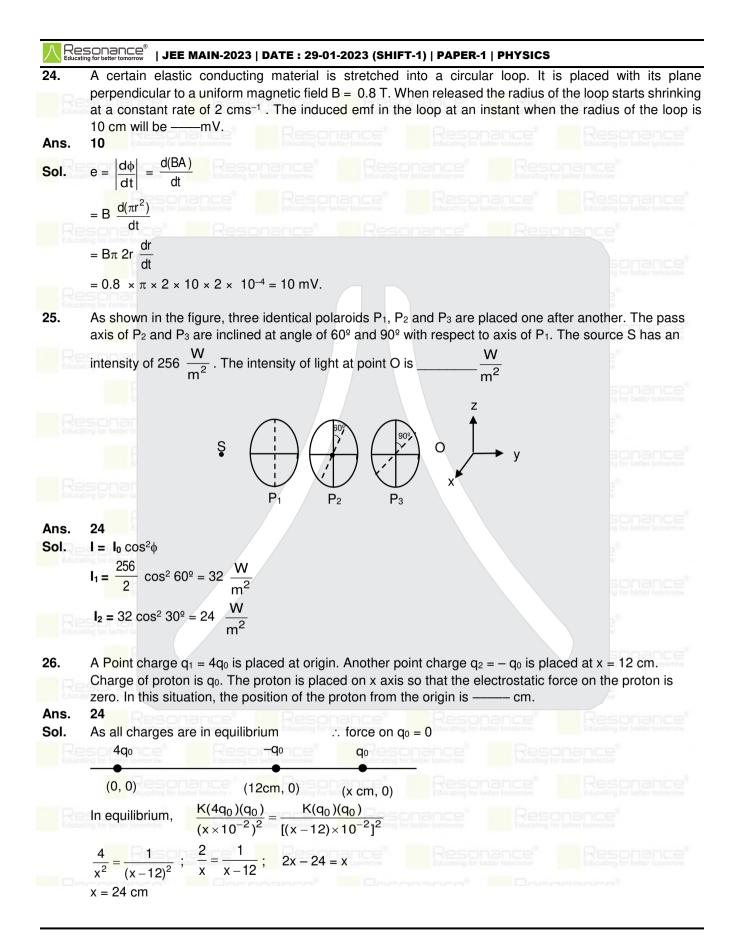
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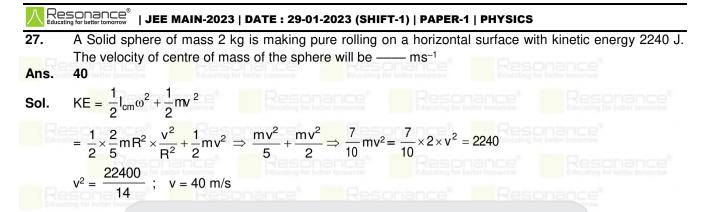
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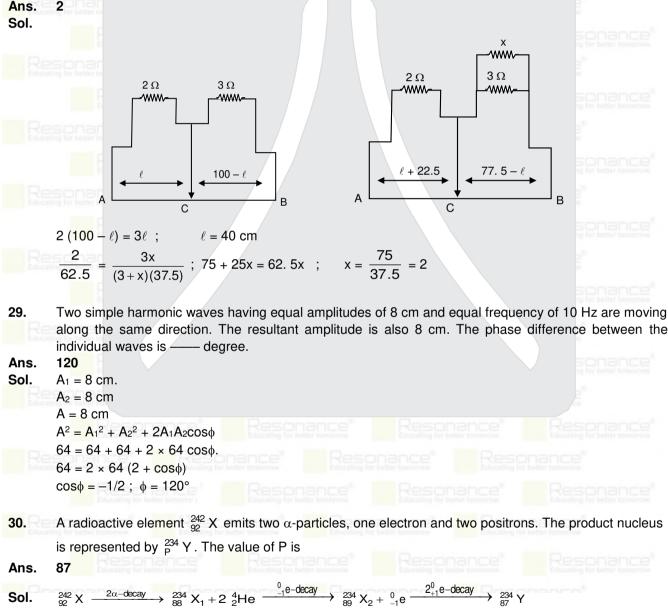


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28. In a metre bridge experiment the balance point is obtained if the gaps are closed by 2Ω and 3Ω . A shunt of X Ω is added to 3Ω resistor to shift the balancing point by 22.5 cm. The value of X is _____.



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P = 87

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