



NEET 2023

Questions, Answer Key & Solutions

Date: 07 May, 2023 | TIME: (02:00 PM to 05:20 PM)

Duration: 200 minutes (03 Hrs. 20 Min.) | Max. Marks: 720

Important Instructions:

- The Answer Sheet is inside this Test Booklet. When you are directed to open the Test Booklet, take out the Answer Sheet and fill in the particulars on OFFICE Copy carefully with **blue/black** ball point pen only.
- The test is of **3 hours 20 minutes** duration and Test Booklet contains **200** multiple-choice questions (four options with a single correct answer) from **Chemistry, Physics and Biology (Botany and Zoology)**. **50** questions in each subject are divided into **two Sections (A and B)** as per details given below:
 - Section A** shall consist of **35 (Thirty-five)** Questions in each subject (Questions Nos – 1 to 35, 51 to 85, 101 to 135 and 151 to 185). All questions are compulsory.
 - Section B** shall consist of **15 (Fifteen)** questions in each subject (Question Nos - 36 to 50, 86 to 100, 136 to 150 and 186 to 200). In Section B, a candidate needs to **attempt any 10 (Ten) questions out of 15 (Fifteen)** in each subject.
Candidates are advised to read all 15 questions in each subject of Section B before they start attempting the question paper. In the event of a candidate attempting more than ten questions, **the first ten questions answered by the candidate shall be evaluated.**
- Each question carries **4 marks**. For each correct response, the candidate will get 4 marks. For each incorrect response, **one mark** will be deducted from the total scores. **The maximum marks are 720.**
- Use **Blue/Black Ball Point Pen only** for writing particulars on this page/markings responses on Answer Sheet.
- Rough work is to be done on the space provided for this purpose in the Test Booklet only.
- On completion of the test, the candidate **must hand over the Answer Sheet (ORIGINAL and OFFICE Copy) to the Invigilator** before leaving the Room/Hall. The candidates are allowed to take away this Test Booklet with them.
- The CODE for this Booklet is F6. Make sure that the CODE printed on the Original Copy of the Answer Sheet is the same as that on this Test Booklet.** In case of discrepancy, the candidate should immediately report the matter to the Invigilator for replacement of both the Test Booklet and the Answer Sheet.
- The candidates should ensure that the Answer Sheet is not folded. Do not make any stray marks on the Answer Sheet. Do not write your Roll No. anywhere else except in the specified space in the Test Booklet/Answer Sheet.
- Use of white fluid for correction is **NOT** permissible on the Answer Sheet.
- Each candidate must show on-demand his/her Admit Card to the Invigilator.
- No candidate, without special permission of the centre Superintendent or Invigilator, would leave his/her seat.
- The candidates should not leave the Examination Hall without handing over their Answer Sheet to the Invigilator on duty and sign (with time) the Attendance Sheet twice. **Cases, where a candidate has not signed the Attendance Sheet second time, will be deemed not to have handed over the Answer Sheet and dealt with as an Unfair Means case.**
- Use of Electronic/ Manual Calculator is prohibited.
- The candidates are governed by all Rules and Regulations of the examination with regard to their conduct in the Examination Room/Hall. All cases of unfair means will be dealt with as per the Rules and Regulations of this examination.
- No part of the Test Booklet and Answer Sheet shall be detached under any circumstances.**
- The candidates will write the Correct Test Booklet Code as given in the Test Booklet/Answer Sheet in the Attendance Sheet.
- Compensatory time of one hour five minutes will be provided for the examination of three hours and 20 minutes duration, whether such candidate (having a physical limitation to write) uses the facility of scribe or not.

In case of any ambiguity in translation of any question, English version shall be treated as final.

प्रश्नों के अनुवाद में किसी अस्पष्टता की स्थिति में, अंग्रेजी संस्करण को ही अन्तिम माना जायेगा।

Name of the Candidate (in Capital letters): _____

Roll Number: in figures:

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 in words: _____

Name of Examination Centre (in Capital letters): _____

Candidate's Signature: _____

Invigilator's Signature: _____

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PART : CHEMISTRY

51. Given below are two statements: one is labelled as **Assertion A** and the other is labelled as **Reason R**:
Assertion A: Metallic sodium dissolves in liquid ammonia giving a deep blue solution, which is paramagnetic.

Reasons R: The deep blue solution is due to the formation of amide.

In the light of the above statements, choose the correct answer from the options given below:

(1) Both **A** and **R** are true but **R** is **NOT** the correct explanation of **A**.

(2) **A** is true but **R** is false.

(3) **A** is false but **R** is true.

(4) Both **A** and **R** are true and **R** is the correct explanation of **A**.

नीचे दो कथन दिए गए हैं। एक को 'अभिकथन A' और दूसरे को 'कारण R' चिन्हित किया गया है।

अभिकथन A : धात्विक सोडियम द्रव अमोनिया में घुलकर गहरे नीले रंग का विलयन देता है, जो अनुचुंबकीय होता है।

कारण R : गहरा नीला विलयन ऐमाइड के बनने के कारण होता है।

ऊपर दिए गए कथनों के आधार पर, नीचे दिए गए विकल्पों में से सही उत्तर चुनिए :

(1) A और R दोनों सत्य हैं, परन्तु R, A की सही व्याख्या नहीं है।

(2) A सत्य है परन्तु R असत्य है।

(3) A असत्य है परन्तु R सत्य है।

(4) A और R दोनों सत्य हैं और R, A की सही व्याख्या है।

Ans. (2)

Sol. A is true but R is false

The deep blue solutions is due to ammoniated electron.



Blue colour is due to ammoniated electron Paramagnetic Nature is due to Ammoniated electron.

52. The conductivity of centimolar solution of KCl at 25°C is 0.0210 ohm⁻¹ cm⁻¹ and the resistance of the cell containing the solution at 25°C is 60 ohm. The value of cell constant is -

(1) 3.28 cm⁻¹ (2) 1.26 cm⁻¹ (3) 3.34 cm⁻¹ (4) 1.34 cm⁻¹

25°C पर KCl के सेंटीमोलर विलयन की चालकता 0.0210 ohm⁻¹ cm⁻¹ है और 25°C पर विलयन वाले सेल का प्रतिरोध 60 ohm है। सेल स्थिरांक का मान है :

(1) 3.28 cm⁻¹ (2) 1.26 cm⁻¹ (3) 3.34 cm⁻¹ (4) 1.34 cm⁻¹

Ans. (2)

Sol. For KCl Solution

$$K = 0.0210 \text{ ohm}^{-1} \text{ cm}^{-1}$$

$$R = 60 \Omega$$

$$\text{conductivity (K)} = \left(\frac{l}{a}\right) \frac{1}{R}$$

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$$\text{cell constant } \left(\frac{\ell}{a}\right) = (K) R$$

$$= 0.0210 \times 60$$

$$= 1.26 \text{ cm}^{-1}$$

53. For a certain reaction, the rate = $k[A]^2[B]$, when the initial concentration of A is tripled keeping concentration of B constant, the initial rate would

- (1) increase by a factor of six.
- (2) increase by a factor of nine.
- (3) increase by a factor of three.
- (4) decrease by a factor of nine.

Ans. (2)

Sol. Rate = $k[A]^2[B]$

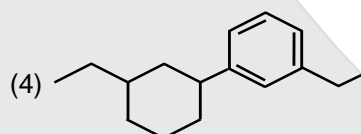
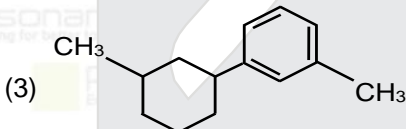
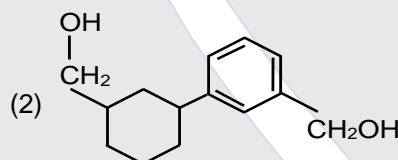
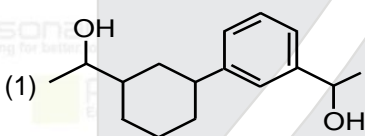
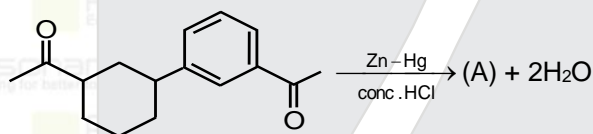
$$r_1 = k(a)^2(b)$$

$$r_2 = k(3a)^2(b)$$

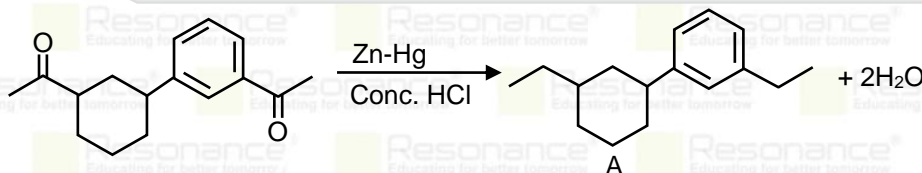
$$\frac{r_2}{r_1} = 9$$

$$r_2 = 9r_1$$

54. Identify product (A) in the following reaction:



Ans. (4)



Sol.

55. Which one is an example of heterogenous catalysis?

- (1) Hydrolysis of sugar catalysed by H^+ ions.
- (2) Decomposition of ozone in presence of nitrogen monoxide.
- (3) Combination between dinitrogen and dihydrogen to form ammonia in the presence of finely divided iron.

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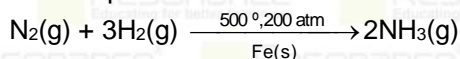
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(4) Oxidation of sulphur dioxide into sulphur trioxide in the presence of oxides of nitrogen.

Ans. (3)

Sol. In heterogeneous catalysis physical state of reactant and catalyst is different.

Haber's process



56. Given below are two statements: one is labelled as **Assertion A** and the other is labelled as **Reason R**:

Assertion A: Helium is used to dilute oxygen in diving apparatus.

Reasons R: Helium has high solubility in O₂. In the light of the above statements, choose the **correct** answer from the options given below:

(1) Both **A** and **R** are true and **R** is **NOT** the correct explanation of **A**.

(2) **A** is true but **R** is false.

(3) **A** is false but **R** is true.

(4) Both **A** and **R** are true and **R** is the correct explanation of **A**.

Ans. (2)

Sol. Helium is used to dilute oxygen in diving apparatus because at high pressure Helium have low solubility in Blood.

57. Amongst the following, the total number of species NOT having eight electrons around central atom in its outer most shell, is NH₃, AlCl₃, BeCl₂, CCl₄, PCl₅

(1) 2

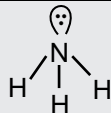
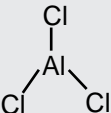
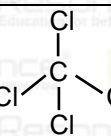
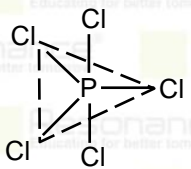
(2) 4

(3) 1

(4) 3

Ans. (3)

Sol.






Species	Structure	Total electron around central atom in outer most shell
NH ₃		8
AlCl ₃		6
BeCl ₂	Cl - Be - Cl	4
CCl ₄		8
PCl ₅		10

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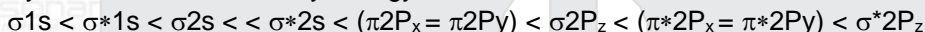
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58. The correct order of energies of molecular orbitals of N_2 molecule, is:

- (1) $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < \sigma 2p_z < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y) < \sigma^* 2p_z$
- (2) $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < \sigma 2p_z < \sigma^* 2p_z < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y)$
- (3) $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y) < \sigma 2p_z < \sigma^* 2p_z$
- (4) $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < (\pi 2p_x = \pi 2p_y) < \sigma 2p_z < (\pi^* 2p_x = \pi^* 2p_y) < \sigma^* 2p_z$

Ans. (4)

Sol. By Molecular orbital theory Energy order of N_2 molecule is



59. Match List-I with List - II :

List - I

- A. Coke
B. Diamond
C. Fullerene
D. Graphite

Choose the correct answer from the options given below :

- (1) A-IV, B-I, C-II, D-III (2) A-III, B-I, C-IV, D-II (3) A-III, B-IV, C-I, D-II (4) A-II, B-IV, C-I, D-III

सूची - I का सूची - II के साथ मिलान कीजिए।

सूची - I

- A. कोक
B. हीरा
C. फुलरीन
D. ग्रेफाइट

सम

- (1) A-IV, B-I, C-II, D-III (2) A-III, B-I, C-IV, D-II (3) A-III, B-IV, C-I, D-II (4) A-II, B-IV, C-I, D-III

Ans. (2)

Sol. List - I

- (A) Coke
(B) Diamond
(C) Fullerene
(D) Graphite

List-II

- I. used as reducing agent
II. carbon atoms are sp^3 Hybridised
III. Cage like molecules
IV. used as dry lubricant

List - II

- I. Carbon atoms are sp^3 hybridised.
II. Used as a dry lubricant
III. Used as a reducing agent
IV. Cage like molecules

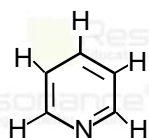
सूची - II

- I. कार्बन परमाणु sp^3 संकरित होते हैं।
II. शुष्क स्नेहक के रूप में उपयोग किया जाता है।
III. अपचायक की भाँति उपयोग किया जाता है।
IV. पिंजरानुमा अणु

60. The number of σ bonds, π bonds and lone pair of electrons in pyridine, respectively are:

- (1) 12, 3, 0 (2) 11, 3, 1 (3) 12, 2, 1 (4) 11, 2, 0

Ans. (2)



Sol. Pyridine

σ Bonds = 11

π Bonds = 3

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Lone pair = 1

61. The element expected to form largest ion to achieve the nearest noble gas configuration is :

- (1) F (2) N (3) Na (4) O

Ans. (2)

Sol. Element For Noble gas electronic configuration

F	F ⁻
N	N ³⁻
Na	Na ⁺
O	O ²⁻

order of ionic size \Rightarrow N³⁻ > O²⁻ > F⁻ > Na⁺

62. Given below are two statements: one is labelled as **Assertion A** and the other is labelled as Reason R :
Assertion A : A reaction can have zero activation energy.

Reasons R : The minimum extra amount of energy absorbed by reactant molecules so that their energy becomes equal to threshold value, is called activation energy.

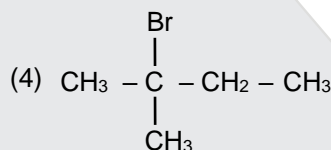
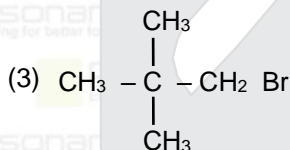
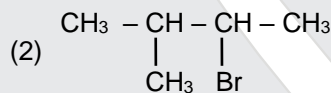
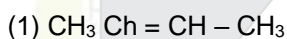
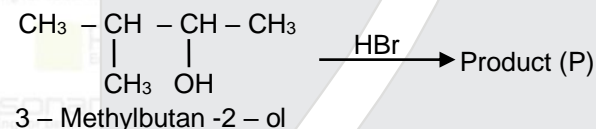
In the light of the above statements, choose the **correct** answer from the option given below :

- (1) Both **A** and **R** are true and **R** is **NOT** the correct explanation of **A**.
 (2) **A** is true but **R** is false.
 (3) **A** is false but **R** is true.
 (4) Both **A** and **R** are true and **R** is the correct explanation of **A**.

Ans. (1)

Sol. Radioactive substance have zero activation energy. So Assertion is true, but reason is not correct explanation because it simple give definition of activation energy.

63. Consider the following reaction and identify the product (P).



Ans. (4)

Sol.

Since, halide is present on allylic position hence, this is an example of allylic halide



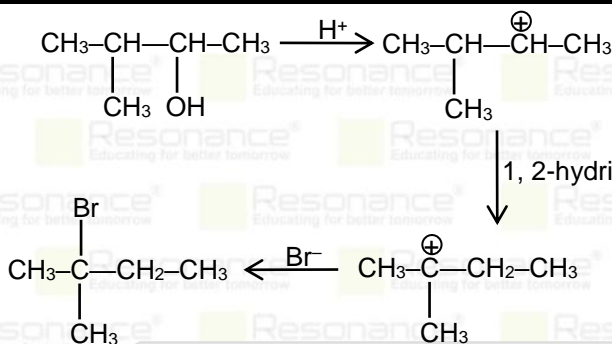
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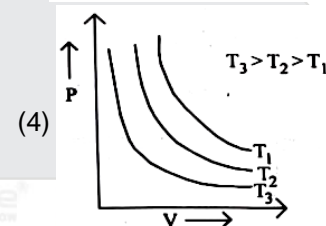
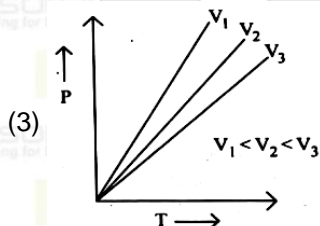
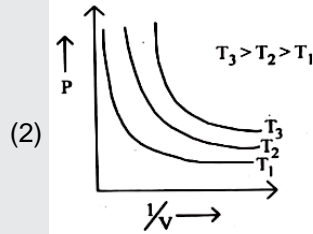
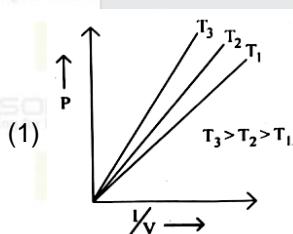


64. Given below are two statements: one is labelled as **Assertion A** and the other is labelled as **Reason R** :
Assertion A : In equation $\Delta_r G = -nFE_{\text{cell}}$, value of $\Delta_r G$ depends on n .
Reasons R : E_{cell} is an intensive property and $\Delta_r G$ is an extensive property.
 In the light of the above statements, choose the correct explanation of answer from the options given below :
- (1) Both **A** and **R** are true and **R** is **NOT** the correct explanation of **A**.
 - (2) **A** is true but **R** is false
 - (3) **A** is false but **R** is true.
 - (4) Both **A** and **R** are true and **R** is the correct explanation of **A**.

Ans. (4)

Sol. $\Delta G = -nF E_{\text{cell}}$
 value of ΔG depends on value of n and because ΔG is an extensive property while E_{cell} is an intensive property

65. Which amongst the following options is correct graphical representation of Boyle's Law?



Ans. (1)

Sol. Boyles' Law

$$P \propto \frac{1}{V} \quad [\text{at constant Temperature and constant amount of gas.}]$$

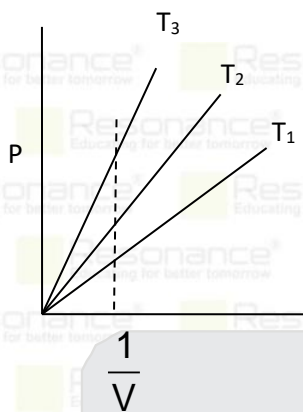
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$$PV = nRT$$

for same value of $\frac{1}{V}$ higher the value of P, higher is temperature. So order of temperature.

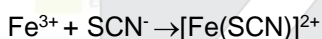
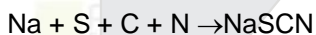
$$\Rightarrow T_3 > T_2 > T_1$$

66. In Lassaigne's extract of an organic compound, both nitrogen and sulphur are present. Which gives blood red colour with Fe^{3+} due to the formation of -

- (1) NaSCN
- (2) $[\text{Fe}(\text{CN})_5\text{NOS}]^{4-}$
- (3) $[\text{Fe}(\text{SCN})]^{2+}$
- (4) $\text{Fe}_4[\text{Fe}(\text{CN})_6]_3 \cdot x\text{H}_2\text{O}$

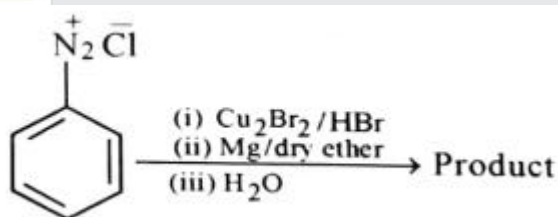
Ans. (3)

Sol. In lassaingne 's extract , when both Nitrogen and Sulphur are present, blood red colour is formed with Fe^{3+} due to formation of $[\text{Fe}(\text{SCN})]^{2+}$



(Blood red colour)

67. Identify the product in the following reaction:



- (1)
- (2)
- (3)
- (4)

Ans. (1)

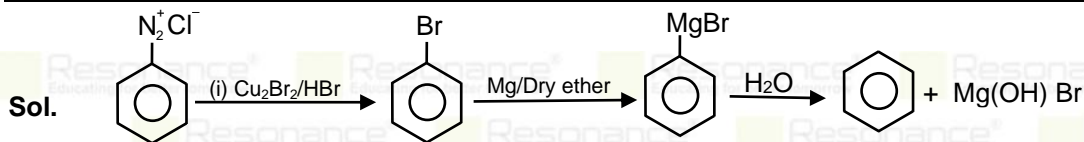
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68. Select the correct **statements** from the following
- Atoms of all elements are composed of two fundamental particles.
 - The mass of the electron is 9.10939×10^{-31} kg.
 - All the isotopes of a given element show same chemical properties
 - Protons and electrons are collectively known as nucleons.
 - Dalton's atomic theory, regarded the atom as an ultimate particle of matter.

- C, D and E only
- A and E only
- B, C and E only
- A, B and C only

Ans. (3)

- Sol. (A) Atoms of all elements are composed of three fundamental Particles
 (B) Mass of electron is 9.1×10^{-31} kg.
 (C) All isotopes of a given element show same chemical properties
 (D) Proton and Neutron are collectively known as nucleors
 (E) Dalton's atomic theory, regard the atom as an ultimate particle of matter
 correct option B, C, E

69. A compound is formed by two elements A and B. The element B forms cubic close packed structure and atoms of A occupy $1/3$ of tetrahedral voids. If the formula of the compound is A_xB_y , the then value of x+y is in options

- 4
- 3
- 2
- 5

Ans. (4)

- Sol. Effective number of atoms B in a unit cell = 4 [CCP]
 Effective number of atoms A in a unit cell = $8 [TV] \times \frac{1}{3} = \frac{8}{3}$

Formula of compound = $A_{\frac{8}{3}} B_4 \Rightarrow A_2B_3$

$A_xB_y = A_2B_3$; so, $x + y = 5$

70. Given below are two statements:

Statement I : A unit formed by the attachment of a base to 1' position of sugar is known as nucleoside

Statement II : When nucleoside is linked to phosphorous acid at 5' -position of sugar moiety, we get nucleotide.

In the light of the above statements, choose the **correct** answer from the options given below :

- Both **Statement I** and **Statement II** are false.
- Statement I** is true but **Statement II** are false.
- Statement I** is false but **Statement II** is true.
- Both **Statement I** and **Statement II** are true.

Ans. (4)

Sol.

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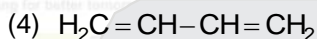
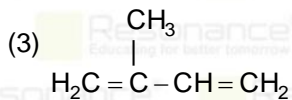
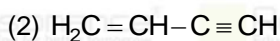
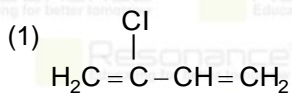
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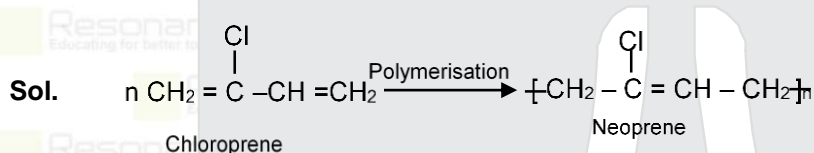
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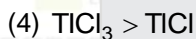
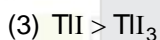
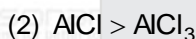
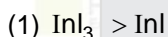
71. Which amongst the following molecules on polymerization produces neoprene?



Ans. (1)



72. Taking stability as the factor, which one of the following represents correct relationship?



Ans. (3)

Sol. due to Inert pair effect in TlI (+1) Oxidation State is more stable than (+3) Oxidation State in TlI₃.

73. Some tranquilizers are listed below. Which one from the following belongs to barbiturates?

(1) Meprobamate

(2) Valium

(3) Veronal

(4) Chlordiazepoxide

Ans. (3)





Sol. veronal is the barbiturates (fact based on NCERT page no. 445)

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74. Which of the following statements are **NOT** correct?
- Hydrogen is used to reduce heavy metal oxides to metals.
 - Heavy water is used to study reactions mechanism.
 - Hydrogen is used to make saturated fats from oils
 - The H-H bond dissociation enthalpy is lowest as compared to single bond between two atoms of any element.
 - Hydrogen reduces oxides of metal that are more active than iron.
- Choose the most appropriate answer from the options given below.

- B, D only
- D, E only
- A, B, C only
- B, C, D, E only

Ans. (2)

Sol.	(D)	Molecule	Bond dissociation enthalpy
		H ₂	435.88 KJ/Mole
		F ₂	158.8 KJ/Mole

The H-H Bond dissociation enthalpy of dihydrogen is highest for single bond between two atom of any element.

(E) Hydrogen can not reduce oxides of metal that are more active than iron.
so statement D & F are incorrect.

75. Intermolecular forces are forces of attraction and repulsion between interacting particles that will include
- dipole – dipole forces.
 - dipole – Induced dipole forces.
 - hydrogen bonding.
 - covalent bonding.
 - dispersion forces

Choose the most appropriate answer form the options given below:

- A, B, C, D are correct.
- A, B, C, E are correct.
- A, C, D, E are correct.
- B, C, D, E are correct.

Ans. (2)

Sol. Covalent band is not intermolecular force of attraction it is band between two atom.






Extra Apart from covalent bonding all the other given inter molecular forces will be included for interacting particles.

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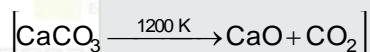
76. Amongst the given options which of the following molecules/ ion acts as a Lewis acid?

- (1) H₂O
(2) BF₃
(3) OH⁻
(4) NH₃

Ans. (2)

Sol. BF₃ act as Lewis acid it has vacant orbitls (6 electron on central atom)

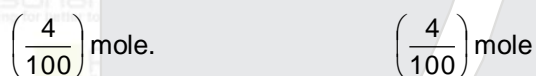
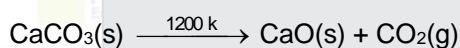
77. The **right** option for the mass of CO₂ produced by heating 20 g of 20% pure limestone is (Atomic mass of Ca = 40)



- (1) 1.76 g (2) 2.64 g (3) 1.32 g (4) 1.12g

Ans. (1)

Sol. Mass of pure line stone (CaCO₃) = $\frac{20 \times 20}{100} = 4$ gram



$$\text{Mass of CO}_2 = \frac{4}{100} \times 44 = 1.76 \text{ gram}$$

78. The relations between n_m. (n_m. = the number of permissible values of magnetic quantum number (m)) for a given value of azimuthal quantum number (l)

- (1) l = 2n_m + 1
(2) n_m = 2l + 1
(3) n_m = l + 2
(4) l = $\frac{n_m - 1}{2}$

Ans. (4)

Sol. Total Passible value of nm = (2l+1)


$$\text{So } l = \left(\frac{n_m - 1}{2}\right)$$

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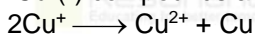
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79. The stability of Cu^{2+} is more than Cu^+ salts in aqueous solution due to -

- (1) enthalpy of atomization.
- (2) hydration energy
- (3) second ionisation enthalpy
- (4) first ionisation enthalpy

Ans. (2)

Sol. Cu (I) compounds are unstable in aqueous solution and undergoes disproportion



The stability of Cu^{2+} (aq) rather than Cu^+ (aq) is due to the much more negative $\Delta H^{\circ}_{\text{Hyd.}}$ of Cu^{2+} (aq) than Cu^+ (aq), which is more than compensates for the second ionisation enthalpy of Cu.

80. Which one of the following statements is correct?

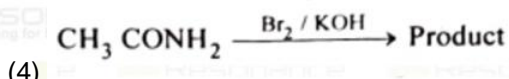
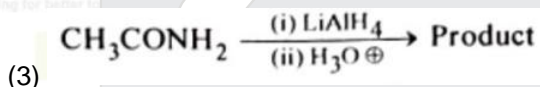
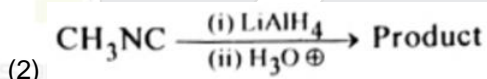
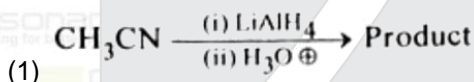
- (1) All enzymes that utilise ATP in phosphate transfer require Ca are the cofactor.
- (2) The bone in human body is an inert and unchanging substance.
- (3) Mg plays roles in neuromuscular function and interneuronal transmission
- (4) The daily requirement of Mg and Ca in the human body is estimated to be 0.2 – 0.3 g

Ans. (4)

Sol. The daily requirement in the human body has been estimated to be 200-300 mg.

- (1) All enzymes that utilize ATP in phosphate transfer require magnesium as the co factor.
- (2) The bone in human body is not inert and unchanging substance.
- (3) Ca plays roles in neuromuscular function and interneuronal transmission.

81. Which of the following reactions will NOT give primary amine as the product?



Ans. (2)

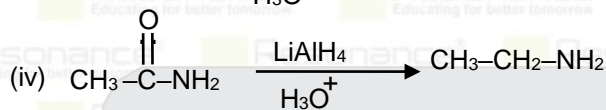
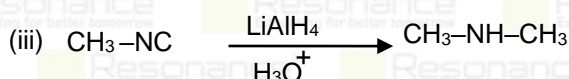
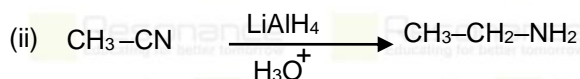
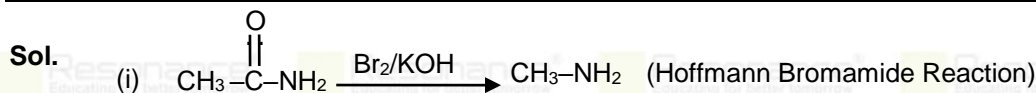
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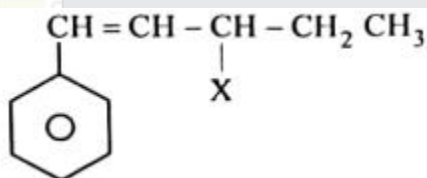
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(iii) Is forming 2° Amine

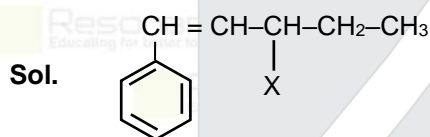
82. The given compound (O)



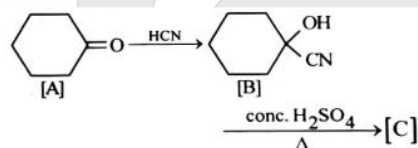
is an example of _____

- (1) aryl halide
- (2) allylic halide
- (3) vinylic halide
- (4) benzylic halide

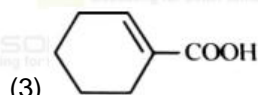
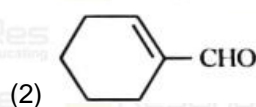
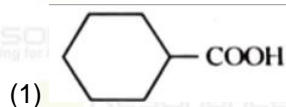
Ans. (2)



83. Complete the following reaction:



(C) is _____



Ans. (3)



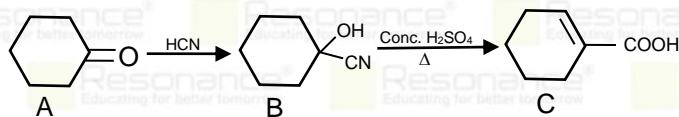
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Dehydration followed by hydrolysis of cyanide group

84. Homoleptic complex from the following complexes is:

- (1) Diamminechloridonitrito – N – platinum (II)
- (2) Pentaamminecarbonatocobalt (III) chloride
- (3) Trimminetriaquachromium (III) chloride
- (4) Potassium trioxalatoaluminate (III)

Ans. (4)

Sol.

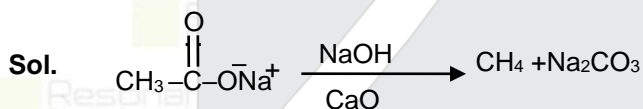
Complex	Nature of complex
(1) $[\text{Pt}(\text{NH}_3)_2\text{Cl}(\text{NO}_2)]$	Heteroleptic
(2) $[\text{Co}(\text{NH}_3)_5(\text{CO}_3)]\text{Cl}$	Heteroleptic
(3) $[\text{Cr}(\text{NH}_3)_5(\text{H}_2\text{O})_3]\text{Cl}_3$	Heteroleptic
(4) $\text{K}_3[\text{Al}(\text{C}_2\text{O}_4)_3]$	Homoleptic

Homoleptic complexes are the complexes in which all ligands are identical : $\text{K}_3[\text{Al}(\text{C}_2\text{O}_4)_3]$

85. Weight (g) of two moles of the organic compound, which is obtained by heating sodium ethanoate with sodium hydroxide in presence of calcium oxide is:

- (1) 32
- (2) 30
- (3) 18
- (4) 16

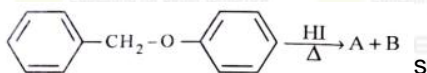
Ans. (1)



Mass of 1 mole of $\text{CH}_4 = 16\text{gm}$

Hence, mass of 2 mole of $\text{CH}_4 = 32\text{gm}$

86. Consider the following reaction :



Identify products A and B.

- (1) A = $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$ and B = $\text{C}_6\text{H}_5\text{I}$
- (2) A = $\text{C}_6\text{H}_5\text{CH}_2\text{I}$ and B = $\text{C}_6\text{H}_5\text{OH}$
- (3) A = $\text{C}_6\text{H}_5\text{CH}_3$ and B = $\text{C}_6\text{H}_5\text{I}$
- (4) A = $\text{C}_6\text{H}_5\text{CH}_3$ and B = $\text{C}_6\text{H}_5\text{OH}$

Ans. (2)

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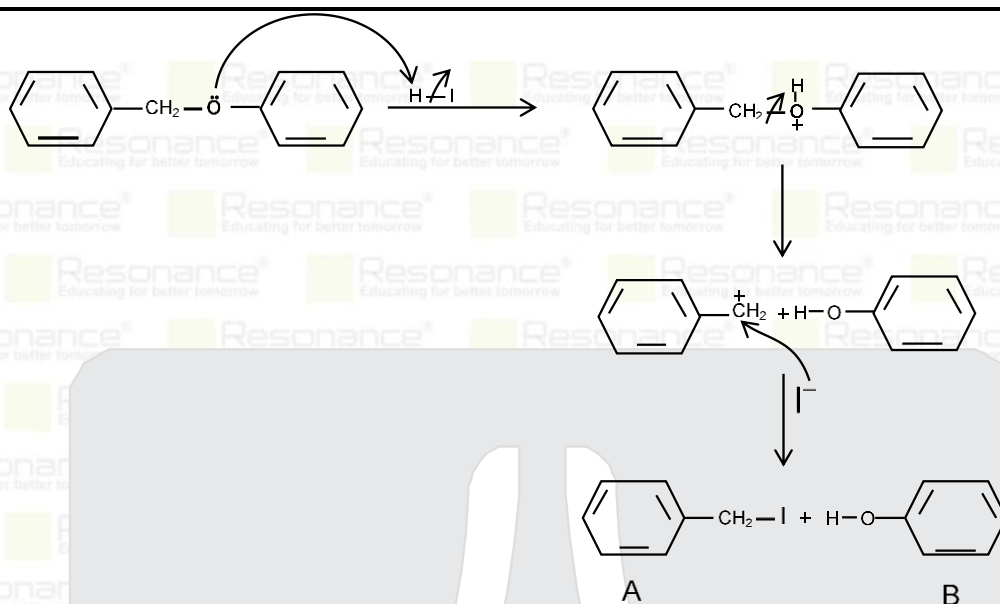
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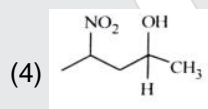
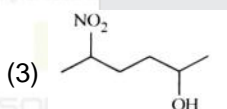
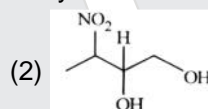
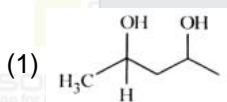
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Sol.



87. Which amongst the following will be most readily dehydrated under acidic conditions?



Ans. (2)

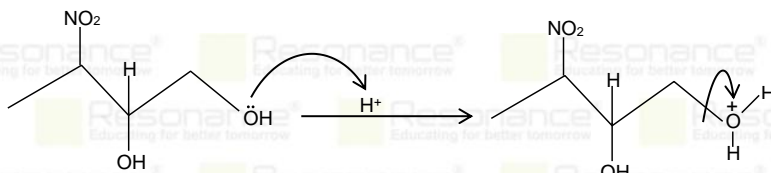
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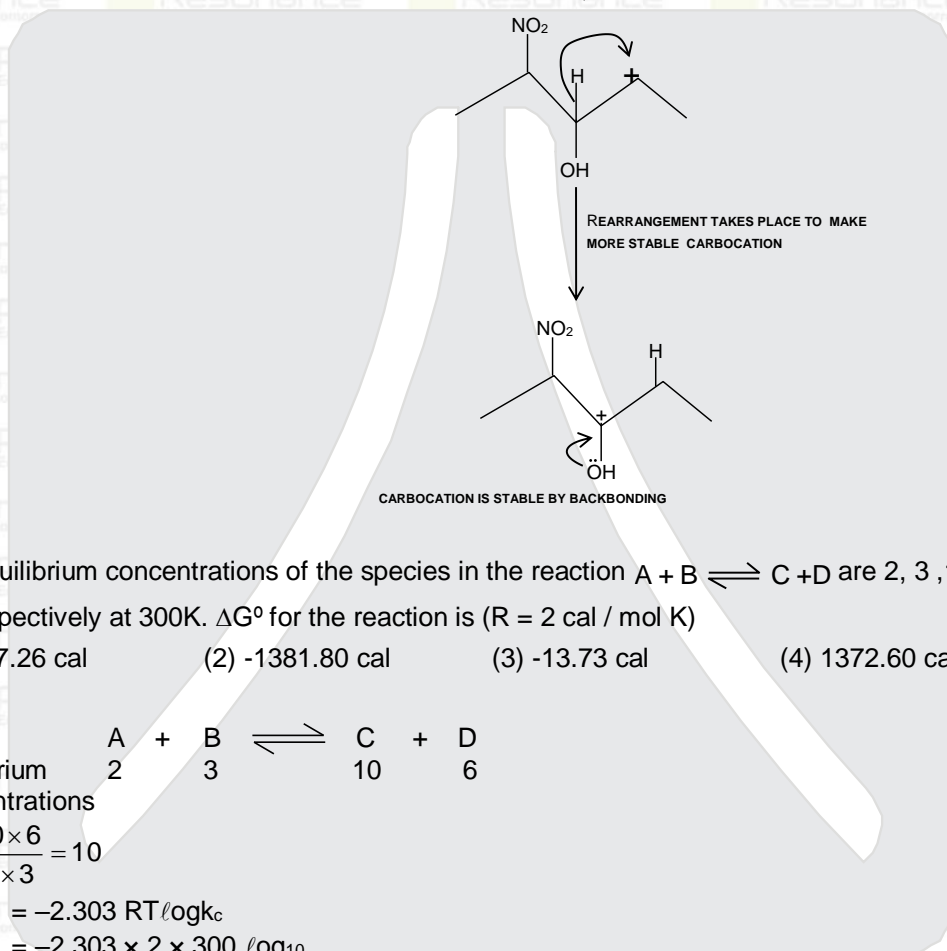
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Sol.



88. The equilibrium concentrations of the species in the reaction $A + B \rightleftharpoons C + D$ are 2, 3, 10 and 6 mol L^{-1} , respectively at 300K. ΔG° for the reaction is ($R = 2 \text{ cal / mol K}$)
- (1) -137.26 cal (2) -1381.80 cal (3) -13.73 cal (4) 1372.60 cal

Ans. (2)

Sol.



$$K_c = \frac{10 \times 6}{2 \times 3} = 10$$

$$\begin{aligned} \Delta G^\circ &= -2.303 RT \log K_c \\ &= -2.303 \times 2 \times 300 \log_{10} 10 \\ &= -1381.80 \text{ cal} \end{aligned}$$

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89. Given below are two statements:

Statement I: The nutrient deficient water bodies lead to eutrophication.

Statement II: Eutrophication leads to decrease in the level of oxygen in the water bodies. In the light of the above statements, choose the correct answer from the options given below:

- (1) Both **Statement I** and **Statement II** are false.
- (2) **Statement I** is correct but **Statement II** is false.
- (3) **Statement I** is incorrect but **Statement II** is true.
- (4) Both **Statement I** and **Statement II** are true.

Ans. (3)

Sol. The correct option is 3

Statement -I is incorrect, the correct explanation is that in eutrophication water bodies nutrient enriched **Statement -II** true

90. Which amongst the following options is the correct relation between change in enthalpy and change in internal energy?

- (1) $\Delta H = \Delta U + \Delta n_g RT$
- (2) $\Delta H - \Delta U = -\Delta n_g RT$
- (3) $\Delta H + \Delta U = \Delta n_g RT$
- (4) $\Delta H = \Delta U - \Delta n_g RT$

Ans. (1)

Sol. Relations between change in enthalpy and change $\Delta H = \Delta U + \Delta n_g RT$

91. Match List I with List - II :

List I (Oxoacids of Sulphur)

- A. Peroxodisulphuric acid
- B. Sulphuric acid
- C. Pyrosulphuric acid
- D. Sulphurous acid

List - II (Bonds)

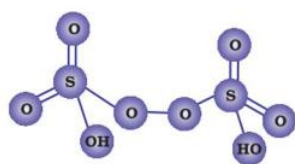
- i. Two S-OH, Four S=O, One S-O-S
- ii. Two S-OH, One S=O
- iii. Two S-OH, Four S=O, One S-O-O-S
- iv. Two S-OH, Two S=O

Choose the correct answer from the options given below:

- (1) A-III, B-IV, C-I, D-II
- (2) A-I, B-III, C-IV, D-II
- (3) A-III, B-IV, C-II, D-I
- (4) A-I, B-III, C-II, D-IV

Ans. (1)

Sol. (A)



Peroxodisulphuric acid
($H_2S_2O_8$)



Sulphuric acid
(H_2SO_4)

(B)

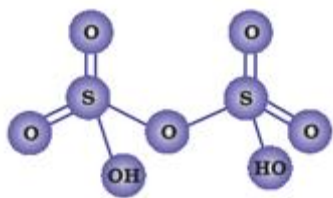
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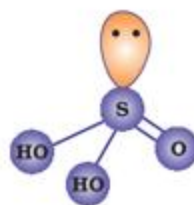
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Pyrosulphuric acid (Oleum)
($H_2S_2O_7$)



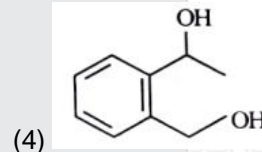
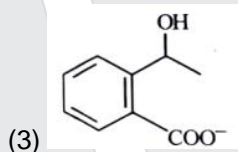
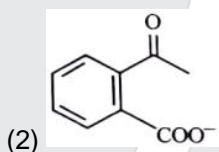
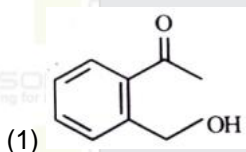
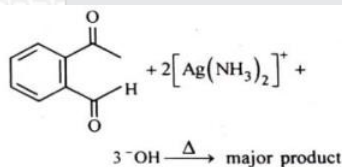
Sulphurous acid
(H_2SO_3)

(C)

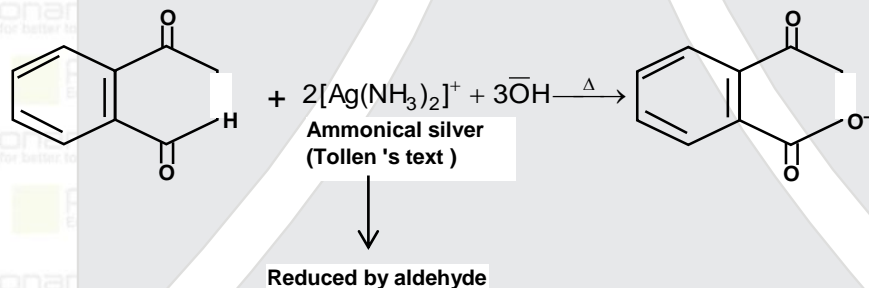
(D)

A → III, B → IV, C → I, D → II

92. Identify the major product obtained in the following reaction:



Ans. (2)



Sol.

So, aldehyde oxidised to carboxylate in basic medium

93. Pumice stone is an example of -

(1) gel

(2) solid sol

(3) foam

(4) sol

Ans. (2)

Sol. Pumice stone.

Example	Disperse phase	Dispersemedium	Type of Colloid
Pumice Stone	Gas	Solid	Solid sol.

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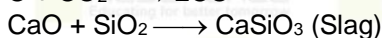
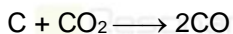
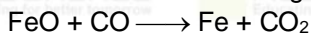
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94. The reaction that does NOT take place in a blast furnace between 900 K to 1500 K temperature range during extraction of iron is:

- (1) $\text{FeO} + \text{CO} \rightarrow \text{Fe} + \text{CO}_2$ (2) $\text{C} + \text{CO}_2 \rightarrow 2\text{CO}$
 (3) $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$ (4) $\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow 2\text{FeO} + \text{CO}_2$

Ans. (4)

Sol. In Blast furnace following reaction are take place in temperature range 900K to 1500K



While reaction at temperature range 500K to 800K



95. Which of the following statements are **INCORRECT**?

- A. All the transition metals except scandium form MO oxides which are ionic.
 B. The highest oxidation number corresponding to the group number in transition metal oxides is attained in Sc_2O_3 to Mn_2O_7 .
 C. Basic character increases from V_2O_3 to V_2O_4 to V_2O_5 .
 D. V_2O_4 dissolves in acids to give VO_4^{3-} salts.
 E. CrO is basic but Cr_2O_3 , is amphoteric.

Choose the correct answer from the options given below:

- (1) B and D only (2) C and D only
 (3) B and C only
 (4) A and E only

Ans. (3)

Sol. Reference NCERT Class-XIIth d =Block (page-No 231

(A) correct statement

(B) Correct statement

(C) $\text{V}_2\text{O}_3 \Rightarrow$ Basic

$\text{V}_2\text{O}_4 \Rightarrow$ Less Basic

$\text{V}_2\text{O}_5 \Rightarrow$ Amphoteric (But mainly acidic)

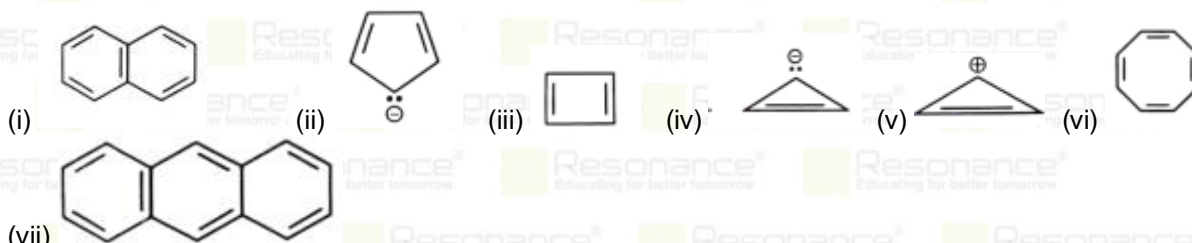
(D) $\text{V}_2\text{O}_4 + 4\text{H}^+ \longrightarrow 2\text{VO}^{2+} + 2\text{H}_2\text{O}$

(E) $\text{CrO} \Rightarrow$ Basic

$\text{Cr}_2\text{O}_3 \Rightarrow$ Amphoteric

So option B and C are in correct.

96. Consider the following compound / species :



The number of compounds / species which obey Huckel's rule is _____

- (1) 6 (2) 2 (3) 5 (4) 4

Ans. (4)






Sol. Huckel's rule

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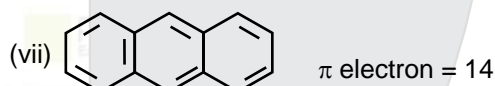
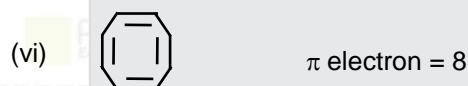
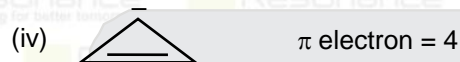
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compound should follow $(4n + 2) \pi$ electron



(i), (ii), (v) & (vii) follows Huckel's rule.

97. What fraction of one edge centred octahedral void lies in one unit cell of fcc?

- (1) $\frac{1}{3}$ (2) $\frac{1}{4}$ (3) $\frac{1}{12}$ (4) $\frac{1}{2}$

Ans. (2)

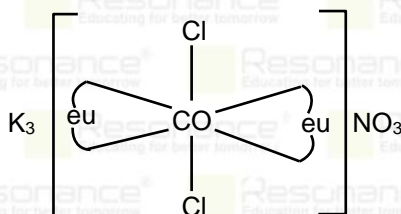
Sol. Contribution of edge centered octahedral void in fcc unit cell = $\frac{1}{4}$

98. Which complex compound is most stable?

- (1) $[\text{Co}(\text{NH}_3)_3(\text{NO}_3)_3]$ (2) $[\text{CoCl}_2(\text{en})_2]\text{NO}_3$
(3) $[\text{Co}(\text{NH}_3)_6]_2(\text{SO}_4)_3$ (4) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Br}](\text{NO}_3)_2$

Ans. (2)

Sol. Because its form stable chelate complex.








Complex which contain chelating ligand have more stability.

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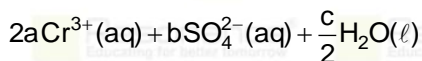
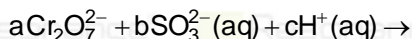
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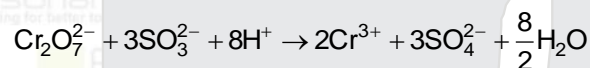
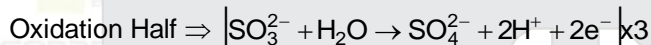
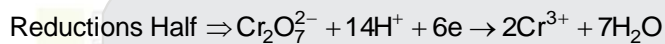
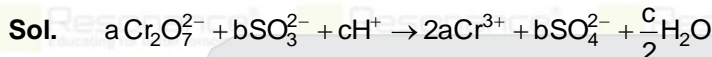
99. On balancing the given redox reaction.



the coefficients a, b and c are found to be, respectively-

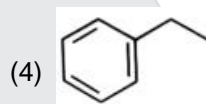
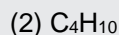
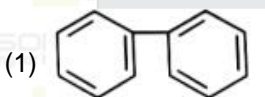
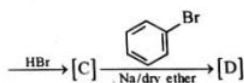
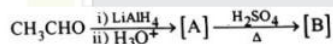
- (1) 3,8,1 (2) 1,8,3 (3) 8,1,3 (4) 1,3,8

Ans. (4)



a = 1 b = 2 c = 8

100. Identify the final product [D] obtained in the following sequence of reactions



Ans. (4)

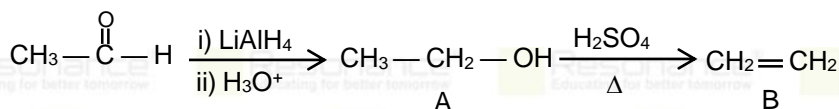
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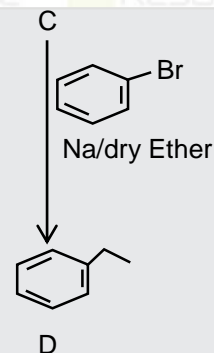
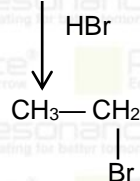
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