



TARGET : NEET (UG) 2024

Course : SARANSH (Youtube Live CRASH COURSE)

I-CHEMISTRY

**DPP**

DAILY PRACTICE PROBLEMS

**DPP NO. 1**

## Organic Chemistry : Periodic Table

### DPP No. : 1

SR. No.	DPPs Qs. Details		Marking Scheme				Time Details			
	Type of Questions	Code	Full Marks	(-ve Marks	Total Ques.	Total Marks	Qs (in Min.) for Each Qs	Time (in Min.)	Max. Time (in Min.)	
1	MULTIPLE CHOICE QUESTION (ONLY ONE CORRECT OPTION)	MCQ	4	-1	10	40	1		10	
	<b>Total</b>				<b>10</b>	<b>40</b>			<b>10</b>	

- The long form of periodic table has
  - (1) Eight horizontal rows and seven vertical columns
  - (2) Seven horizontal rows and eighteen vertical columns
  - (3) Seven horizontal rows and seven vertical columns
  - (4) Eight horizontal rows and eight vertical columns
- All the s-block element of the periodic table are placed in the group.
  - (1) IA and IIA
  - (2) IIIA and IVA
  - (3) B sub groups
  - (4) VA to VIIA
- The element having the electronic configuration  $1s^2 2s^2 2p^6, 3s^2 3p^1$  is
  - (1) A transition element
  - (2) A representative element
  - (3) An inert gas
  - (4) An inner-transition element
- Elements in which 4f orbitals are progressively filled are called as
  - (1) Transition elements
  - (2) Lanthanides
  - (3) Actinides
  - (4) Inert gases
- Atomic number of 15, 33, 51 represents the following family :
  - (1) carbon family
  - (2) nitrogen family
  - (3) oxygen family
  - (4) None
- Li resembles Mg due to diagonal relationship which is attributed to :
  - (1) nearly similar polarising power
  - (2) same value of electron affinity
  - (3) penetration of sub-shells
  - (4) identical effective nuclear charge

7. Which of the following are d-block elements but not transition elements ?  
 (1) Zn, Mn, Fe                      (2) Zn, Cd, Hg                      (3) Y, La, Sc                      (4) Au, Pd, Pt
8. Which one of the following is the biggest ion  
 (1)  $\text{Al}^{+3}$                       (2)  $\text{Ba}^{+2}$                       (3)  $\text{Mg}^{+2}$                       (4)  $\text{Na}^{+}$
9. What is the position of the element in the Modern periodic table satisfying the electronic configuration  $(n - 1) d^1 ns^2$  for  $n = 4$  :  
 (1) 3rd period and 3rd group                      (2) 4th period and 4th group  
 (3) 3rd period and 2nd group                      (4) 4th period and 3rd group
10. Which of the following statement is correct :  
 (1) Due to lanthanide contraction size of 3d series elements  $\approx$  4d series element .  
 (2) Due to lanthanide contraction size of 4d series elements  $\approx$  5d series element.  
 (3) Due to lanthanide contraction size of 3d series elements  $<$  5d series element.  
 (4) Due to lanthanide contraction size of 5d series element  $>$  4d series elements.

### Answer Key

1. (2)    2. (1)    3. (2)    4. (2)    5. (2)    6. (1)    7. (2)  
 8. (2)    9. (4)    10. (2)