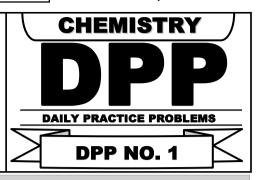
**TARGET: NEET (UG) 2024** 

Course: SARANSH (Youtube Live CRASH COURSE)



## **CHEMISTRY: Redox Reaction**

		<u>DPP N</u>	<u>o. : 1</u>			
1.	Oxidation number of f	luorine in OF <sub>2</sub> is :	(3) –1	(4) –2		
2.	In Ni(CO) <sub>4</sub> , the oxidati	. ,	. ,	( /		
	(1) 4	(2) 0	(3) 2	(4) 8		
3.	Oxidation state of nitro Compound (1) [Co(NH <sub>3</sub> ) <sub>5</sub> Cl]Cl <sub>2</sub> (3) (N <sub>2</sub> H <sub>5</sub> ) <sub>2</sub> SO <sub>4</sub>	ogen is incorrectly giver -3 +2	o for:  Oxidation State  (2) $NH_2OH$ -  (4) $Mg_3N_2$ -			
4.	Oxidation number of F (1) 200	Fe in Fe <sub>0.94</sub> O is : (2) 200/94	(3) 94/200	(4) None		
5.	In the reaction 4Fe + 3O₂ → 4Fe³+ + 6O²- which of the following statements is incorrect?  (1) It is a redox reaction (2) Metallic iron is a reducing agent (3) Fe³+ is an oxidising agent (4) Metallic iron is reduced to Fe³+					
6.	The compound that can work both as an oxidising as well as reducing agent is : (1) $KMnO_4$ (2) $H_2O_2$ (3) $Fe_2(SO_4)_3$ (4) $K_2Cr_2O_7$					
7.	In a reaction of					

 $H_2O$  (steam) + C (glowing)  $\rightarrow$  CO +  $H_2$ 

(1) H<sub>2</sub>O is the reducing agent

(2) H<sub>2</sub>O is the oxidising agent

(3) carbon is the oxidising agent

(4) oxidation-reduction does not occur

 $H_2O$  (steam) + C (glowing)  $\rightarrow$  CO +  $H_2$ 

- 8. If an element is in its lowest oxidation state, under proper conditions it can act as:
  - (1) Reducing agent
  - (2) An oxidising agent
  - (3) Oxidising as well as reducing agent
  - (4) Neither oxidising nor reducing agent
- 9. In the reaction,

 $2S_2O_3^{2-} + I_2 \longrightarrow S_4O_6^{2-} + 2I$ , the eq. wt. of  $S_2O_3^{2-}$  is equal to its :

- (1) Mol. wt.
- (2) Mol. wt./2
- (3) 2 x Mol. wt.

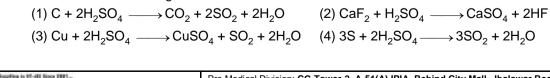
(4) Mol. wt./6



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				SAKANSH   CHEMISTRY			
10.	In the reaction : $A^{-1}$ (1) $n_1 + n_2$	$^{n^2} + xe^- \rightarrow A^{-n_1}$ , here (2) $n_2 - n_1$		(4) n <sub>1</sub> . n <sub>2</sub>			
11.	Equivalent weight (1) M/3	of $FeC_2O_4$ in the change (2) M/6	e: $FeC_2O_4 \rightarrow Fe^{3+} + CO_2$ (3) M/2	<sub>2</sub> is : (4) M/1			
12.	The number of mo (1) 1/5	le of oxalate ions oxidis (2) 2/5	sed by one mole of MnO <sub>2</sub> (3) 5/2	<sub>4</sub> -is: (4) 5			
13.	What will be n-fact (1) 5	or for Ba(MnO <sub>4</sub> ) <sub>2</sub> in acid (2) 10	dic medium? (Where it be	ehaves as oxidant) (4) 3			
14.	Which is a redox reaction : (1) $2CuI_2 \rightarrow 2CuI + I_2$						
	(2) NaCl + AgNO <sub>3</sub> $\rightarrow$ AgCl + NaNO <sub>3</sub> (3) NH <sub>4</sub> Cl + NaOH $\rightarrow$ NH <sub>3</sub> + NaCl + H <sub>2</sub> O (4) $Cr_2(SO_4)_3$ + 6KOH $\rightarrow$ 2Cr(OH) <sub>3</sub> + 3K <sub>2</sub> SO <sub>4</sub>						
15.	Which of the following change represents a disproportionation reaction (s) : (1) $Cl_2 + 2OH^- \rightarrow ClO^- + Cl^- + H_2O$ (2) $Cu_2O + 2H^+ \rightarrow Cu + Cu^{2+} + H_2O$						
	(3) $2HCuCl_2 \xrightarrow{\text{dilution with}} Cu + Cu^{2+} + 4Cl^{-} + 2H^{+}$ (4) All of the above						
16.	$H_2O_2 + H_2O_2 \longrightarrow 2H_2O + O_2$ is an example of disproportionation because - (1) Oxidation number of oxygen only decreases (2) Oxidation number of oxygen only increases (3) Oxidation number of oxygen decreases as well as increases (4) Oxidation number of oxygen neither decreases nor increases $H_2O_2 + H_2O_2 \longrightarrow 2H_2O + O_2$ एक विशमानुपातन का उदाहरण है, क्योंकि-						
17.	What is the value of $n$ in the following equation : $Cr(OH)_4^- + OH^- \longrightarrow CrO_4^{2-} + H_2O + ne$ ?						
	(1) 3	(2) 6	(3) 5	(4) 2			
18.	In the reaction:  MnO <sub>4</sub> <sup>-</sup> + xH <sup>+</sup> + ne <sup>-</sup> What is the value of  (1) 5	$\longrightarrow Mn^{2+} + yH_2O$ of n: (2) 8	(3) 6	(4) 3			
19.	(a) $H_2O_2 + O_3 \rightarrow H_2O_2 + Ag_2O - G_2O_2 + G_2O_2O_2$ Role of hydrogen p	l <sub>2</sub> O + 2O <sub>2</sub> → Ag + H <sub>2</sub> O + O <sub>2</sub> peroxide in the above real and reducing in (b)	eactions is respectively-	a) and oxidizing in (b)			



does not show oxidizing behaviour?

20.

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Hot concentrated sulphuric acid is a moderately strong oxidizing agent. Which of the following reactions